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USER MANUAL

Philips Evergo Manual

Michael Seilmaier



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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web chem 1210 spring 2019 experiment 10.11 part 1 acid base titration abstract the purpose of this experiment is to observe the titration of hydrochloric acid (strong acid) with sodium hydroxide (a strong base) and acetic acid (a

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web goals understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1:1
 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$
 one mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide (NaOH) and oxalic acid ($\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$) and the molar mass of an unknown solid acid using

the molar mass of the standardized sodium hydroxide solution

experiment 3 acid base titration chem 1202 expt 3 title - Aug 04 2022

web acid base reactions can be used to determine concentrations of an acid or a base using stoichiometric calculations and titrimetric analysis the volume of a solution required to react

lab 6 acid base titration chemistry libretexts - Feb 10 2023

web jul 19 2023 introduction titration is an analytical quantitative technique used to determine the concentration of a solute a pH titration is used to determine the concentration of an acid or a base titrations play an important role in determining amount and purity in many manufacturing processes

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pKa of an unknown acid or pKb of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

lab report titration chm1046l volumetric analysis acid base - Apr 12 2023

web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being NaOH reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4$ aq 2NaOH aq $\text{Na}_2\text{C}_2\text{O}_4$ $4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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data calculations and conclusion acid base titration lab - Jun 02 2022

web conclusion in this lab we used titration to explore the concepts of stoichiometry and equivalence points we found the concentration of an unknown substance by mixing 2M HCl with the NaOH of unknown concentration in order to experimentally ascertain the concentration of the NaOH

this is a chemistry lab report on an acid base titration experiment - Dec 28 2021

web this point of titration is referred to as the equivalence point considering the 1:1 stoichiometry of this acid base reaction NaOH aq $\text{C}_6\text{H}_4\text{COOH}$ aq $\text{C}_6\text{H}_4\text{COONa}$ aq H_2O l the point of equivalence is the point of titration when the number of moles of NaOH added is equal to the number of moles of KHP in the

