

10.1

THE NATURE OF GASES

SECTION REVIEW

Objectives

- Describe the motion of gas particles according to the kinetic theory
- Interpret gas pressures in terms of kinetic theory

Key Terms

- | | | |
|--------------------------------|------------------------------|------------------------|
| • kinetic energy | • gas pressure | • atmospheric pressure |
| • kinetic theory | • vacuum | • barometer |
| • partial (P _{part}) | • standard atmospheric (atm) | |

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are highlighted in this section. Write clearly over the completed work a term, short phrase, or sentence.

- The kinetic theory describes the _____ of particles in matter and the forces of attraction between them. The theory assumes that the volume occupied by a gas is mostly _____, that the particles of gas are _____, move _____ of each other, and are in constant _____ motion. Furthermore, the _____ between particles are elastic so that the total _____ remains constant. The pressure of a gas results from the countless collisions of the gas particles with an object. Barometers are used to measure _____ pressure. Standard conditions are defined as a temperature of _____ and a pressure of _____.
1. _____
 2. _____
 3. _____
 4. _____
 5. _____
 6. _____
 7. _____
 8. _____
 9. _____
 10. _____

Part B True-False

Classify each of these statements as true (always true), AT (sometimes true), ST (sometimes true), or F (false).

- _____ 11. Atmospheric pressure is 760 mm Hg.
- _____ 12. The SI unit of pressure is the pascal.
- _____ 13. Atmospheric pressure decreases as you climb a mountain because the air layer that surrounds the earth thins out.

Review Gases Section 4 Answers

Monica M. Tighe, Marg Brown

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author names.

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