

How to Identify the Limiting & Excess Reactants/Reagents

Molar Mass:

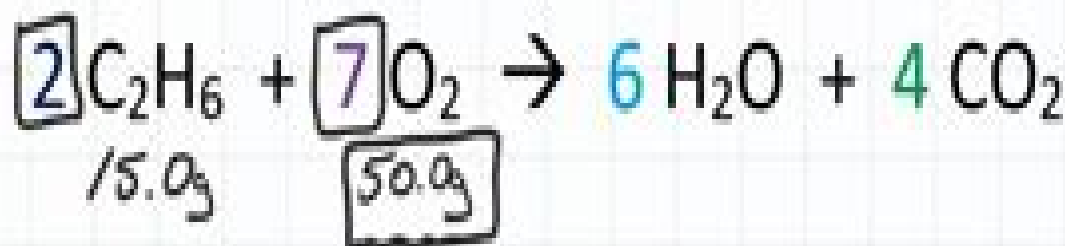
$\text{C}_2\text{H}_6 = 30.08 \text{ g/mol}$

$\text{O}_2 = 32.00 \text{ g/mol}$

$\text{H}_2\text{O} = 18.02 \text{ g/mol}$

1.) Calculate how much of one reactant will react with the other reactant

Ex. What is the limiting reactant when 15.0 g of C_2H_6 react with 50.0 g of O_2 ?



$$\overset{\text{A}}{\text{15.0g } \cancel{\text{C}_2\text{H}_6}} \left(\frac{1 \cancel{\text{mol C}_2\text{H}_6}}{30.08 \cancel{\text{g C}_2\text{H}_6}} \right) \left(\frac{7 \cancel{\text{mol O}_2}}{2 \cancel{\text{mol C}_2\text{H}_6}} \right) \left(\frac{32.00 \cancel{\text{g O}_2}}{1 \cancel{\text{mol O}_2}} \right) = \text{55.9 g } \overset{\text{B}}{\text{O}_2}$$

$\text{O}_2 = \text{L.R.}$

$$\overset{\text{A}}{50.0 \cancel{\text{g O}_2}} \left(\frac{1 \cancel{\text{mol O}_2}}{32.00 \cancel{\text{g O}_2}} \right) \left(\frac{2 \cancel{\text{mol C}_2\text{H}_6}}{7 \cancel{\text{mol O}_2}} \right) \left(\frac{30.08 \cancel{\text{g C}_2\text{H}_6}}{1 \cancel{\text{mol C}_2\text{H}_6}} \right) = 13.4 \text{ g } \overset{\text{B}}{\text{C}_2\text{H}_6}$$

$\text{C}_2\text{H}_6 = \text{E.R.}$

If reacting all of Reactant A requires **more** of Reactant B than is available, then **B** is the Limiting Reactant.

If reacting all of Reactant A requires **less** of Reactant B than is available, then **A** is the Limiting Reactant.

Review Answers Limiting Reactants

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