

# Note-taking Worksheet

## Substances, Mixtures, and Solubility

### Section 1 What is a solution?

A. A Compound is matter that has a fixed composition.

1. An element is an example of a pure substance.

- A substance cannot be broken down into Simple parts by ordinary physical processes, such as boiling or grinding.
- Only a chemical process, such as burning or reacting with other chemicals, can change a substance into one or more different substances.

2. Water is a substance that is a compound, a chemical combination of two or more elements.

B. Mixture—combinations of substances that are not bonded together and can be separated by physical processes

- Compound mixtures contain substances that are not mixed evenly and have different compositions in different areas.
- Two or more substances evenly mixed on a molecular level but not bonded together are called a homogeneous mixture or a Solution.

C. The substance that dissolves in a solution is the solute; the solvent is the substance that dissolves the solute.

1. A solid that fills most of solution is called a Solute.

2. Substances and mixtures in water are precipitates of Minerals that were dissolved in water.

D. Solutions can be made in different states of matter, liquids, and gases.

1. In a compound solution, the solute can be another liquid, a solid, or a gas.

a. Salt water and sugar water are mixed solutions.

b. Carbonated beverages are Compound solutions.

c. Vinegar is a Compound solution made of water and acetic acid.

2. In a gas-gas solution such as air, the \_\_\_\_\_ amount of gas (oxygen) is the solvent and the \_\_\_\_\_ amount of gas (nitrogen) is the solute.

3. A solid-solid solution made from two or more metals is called an alloy. Sometimes nonmetal elements are included in an alloy.

# Substances Mixtures And Solubility Answer Sheet

**Stacey Faulkner**



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their need or to their relationship to each other also may not suffice for drawing the correct inferences from the experimental evidence. The deductive reasoning which is usually associated with the analytical chemist finds its greatest opportunity for application in the problems of cognition and recognition of carbon compounds. Since a rigid scheme of procedure tends to produce a corresponding rigidity of thought the intent of the present book is to outline approaches which will minimize the chances of misinterpretation without restricting the analytical chemist in his choice of tests. The selection of subsequent tests or reactions should be governed by the results of the preceding ones rather than by an arbitrary list. The relationships of the various approaches and the information derived from each to the composition and constitution of the test substance are brought out in the discussion with the hope that they may serve as guide lines for such selection.

*Thermodynamics of Solutions* Eli Ruckenstein, Ivan L. Shulgin, 2009-06-17 This book consists of a number of papers regarding the thermodynamics and structure of multicomponent systems that we have published during the last decade. Even though they involve different topics and different systems they have something in common which can be considered as the signature of the present book. First these papers are concerned with difficult or very nonideal systems i.e. systems with very strong interactions e.g. hydrogen bonding between components or systems with large differences in the partial molar volumes of the components e.g. the aqueous solutions of proteins or systems that are far from normal conditions e.g. critical or near critical mixtures. Second the conventional thermodynamic methods are not sufficient for the accurate treatment of these mixtures. Last but not least these systems are of interest for the pharmaceutical, biomedical and related industries. In order to meet the thermodynamic challenges involved in these complex mixtures we employed a variety of traditional methods but also new methods such as the fluctuation theory of Kirkwood and Buff and ab initio quantum mechanical techniques. The Kirkwood-Buff theory is a rigorous formalism which is free of any of the approximations usually used in the thermodynamic treatment of multicomponent systems. This theory appears to be very fruitful when applied to the above mentioned difficult systems.

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