

FINAL EXAMINATION IN PHYSICAL SCIENCE

I. Multiple Choices: For each of the following items, choose the BEST answer. On the answer sheet provided, shade the bubble that corresponds to the letter of your choice.

- The chemical formula of caffeine is $C_8H_{10}N_4O_2$. What is its molecular mass?
 A. 165 g/mol B. 172 g/mol C. 194 g/mol D. 208 g/mol
- People usually use hydrogen peroxide (H_2O_2 ; molar mass = 34.02 g/mol) to clean their wounds. If Ricardo Delasay used 1.0 g of H_2O_2 to clean his gunshot wound, how many moles of H_2O_2 did he use?
 A. 0.035 mol B. 0.030 mol C. 0.029 mol D. 0.0025 mol
- Methane (CH_4 ; molar mass = 16.04 g/mol) is one of the greenhouse gases. If 32.08 g of CH_4 was released in the atmosphere, how many molecules of CH_4 were added in the atmosphere?
 A. 1.20×10^{24} B. 1.20×10^{23} C. 6.02×10^{24} D. 6.02×10^{23}

For numbers 4 and 5, consider the problem below:

- Iron (III) oxide reacts with carbon monoxide to form iron and carbon dioxide. Which is the balanced chemical equation that represents the reaction?
 A. $Fe_2O_3 + 3CO \rightarrow 2Fe + 3CO_2$ C. $2Fe_2O_3 + 3CO \rightarrow 4Fe + 3CO_2$
 B. $Fe_2O_3 + CO \rightarrow Fe + CO_2$ D. $Fe_2O_3 + 3CO \rightarrow Fe + 3CO_2$
- What mass of iron (III) oxide is needed to produce 5.6 kg of iron?
 A. 4.0 kg B. 8.0 kg C. 12 kg D. 16 kg

For numbers 6-8, consider the problem below:

Ammonium carbonate decomposes to form ammonia gas, carbon dioxide and water. The reaction is represented by the equation below:



- How many moles of ammonia are formed for every mole of the reactant used?
 A. 1 mole B. 2 moles C. 2.5 moles D. 3 moles
- How many moles of the reactant is required to produce 22 grams of CO_2 ?
 A. 0.50 mole B. 1.0 mole C. 1.5 mole D. 2.0 moles
- How many grams of $(NH_4)_2CO_3$ is required to produce 1.5 moles of water?
 A. 49 g B. 96 g C. 144 g D. 192 g

For numbers 9 and 10, consider the problem below:

45.9 g of K_2CO_3 was reacted with an excess of HCl, yielding 48.25 g of KCl.



- What is the theoretical yield?
 A. 46.3 g B. 48.0 g C. 48.3 g D. 49.5 g
- What is the percent yield?
 A. 92.5 % B. 93.5 % C. 97.0 % D. 97.5 %
- Which of the following will cause a soap to reduce its cleaning activity?
 A. Alkyl hydrogen sulfate C. presence of surfactants
 B. Hard water D. soft water
- Which of the following chemicals are linked to breast cancer in women?
 A. Parabens B. PEG C. Paraffin wax D. BHA
- Which of the following disinfectants has a fungicidal property and is a potential carcinogen?

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**Annette Upmeier zu Belzen, Dirk
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