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Answer & Solutions for NEET 2022_(Re-Exam)

Chemistry

51. Match List - I with List - II :

List - I (quantum number)	List- II (orbital)
(a) $n = 2, \ell = 1$	(i) 2 s
(b) $n = 3, \ell = 2$	(ii) 3 s
(c) $n = 3, \ell = 0$	(iii) 2 p
(d) $n = 2, \ell = 0$	(iv) 3 d

Choose the correct answer from the options given below :

- (1) (a) - (iii), (b) - (iv), (c) - (ii), (d) - (i)
 (2) (a) - (iii), (b) - (iv), (c) - (i), (d) - (ii)
 (3) (a) - (iv), (b) - (iii), (c) - (i), (d) - (ii)
 (4) (a) - (iv), (b) - (iii), (c) - (ii), (d) - (i)

Sol. Answer (1)

$\ell = 0 \Rightarrow$ s-subshell

$\ell = 1 \Rightarrow$ p-subshell

$\ell = 2 \Rightarrow$ d-subshell

$\ell = 3 \Rightarrow$ f-subshell

$\therefore n = 2, \ell = 1 \Rightarrow 2 p$

$n = 3, \ell = 2 \Rightarrow 3 d$

$n = 3, \ell = 0 \Rightarrow 3 s$

$n = 2, \ell = 0 \Rightarrow 2 s$

52. The density of the solution is 2.15 g mL^{-1} , then mass of 2.5 mL solution in correct significant figures is :

- (1) 53.75 g (2) $5375 \times 10^{-3} \text{ g}$
 (3) 5.4 g (4) 5.38 g

Sol. Answer (3)

In case of multiplication and division, the final result should be reported as having the same number of significant digits as the number with least number of significant digits.

$$\therefore \text{Density} = \frac{\text{Mass}}{\text{Volume}}$$

$$\begin{aligned}\text{So, Mass} &= 2.15 \times 2.5 \\ &= 5.375 \\ &\approx 5.4 \text{ g}\end{aligned}$$

53. Given below are two statements : one is labelled as Assertion (A) and the other is labelled as Reason (R).

Assertion (A) :

Chlorine is an electron withdrawing group but it is ortho, para directing in electrophilic aromatic substitution.

Reason (R) :

Inductive effect of chlorine destabilises the intermediate carbocation formed during the electrophilic substitution, however due to the more pronounced resonance effect, the halogen stabilises the carbocation at ortho and para positions.

In the light of the above statements, choose the most appropriate answer from the options given below :

- (1) (A) is not correct but (R) is correct.
 (2) Both (A) and (R) are correct and (R) is the correct explanation of (A).
 (3) Both (A) and (R) are correct but (R) is not the correct explanation of (A).
 (4) (A) is correct but (R) is not correct.

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