



## Answers to examination-style questions

Answers	Marks	Examiner's tips
1 a) i) 2	1	
ii) 1	1	
iii) 0	1	It is easier to calculate the order with respect to B first in this example.
b) $k = \frac{\text{rate}}{[D]^2[E]}$  or $\frac{8.36 \times 10^{-4}}{(0.84)^2(1.16)}$ $= 1.02 \times 10^{-3} \text{ mol}^{-2} \text{ dm}^6 \text{ s}^{-1}$	3	Make sure you can rearrange formulae.  The units are worth a mark so do not leave them out.  Answers from $1.02 \times 10^{-3}$ to $1.05 \times 10^{-3}$ are acceptable to take account of rounding up in the numbers during the calculation.
2 a) i) Expt 2 $2.68 \times 10^{-4}$ Expt 3 $10.72 \times 10^{-4}$ Expt 4 $2.08 \times 10^{-3}$	3	Fewer than 2 dp here will be penalised. This gives 3 significant figs in answer.  Accept $10.7 \times 10^{-4}$
ii) $k = \frac{\text{rate}}{[X]^2}$ or $\frac{2.68 \times 10^{-4}}{(1.20 \times 10^{-3})^2}$ $= 186 \text{ mol}^{-1} \text{ dm}^3 \text{ s}^{-1}$	3	Units are worth a mark so don't leave them out.  In this case candidates often do part ii) to calculate $k$ before part i).
b) increases exponentially, i.e. 	1	This is quite hard so a straight line like this was allowed:  but not a curve like this:

# Nelson Thornes Chemistry A2 Answers Chapter 13

**E. N. Ramsden**



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