

Physical Chemistry II – Exam 1 SOLUTIONS

Constant	Value in SI Units
h	$6.62607 \times 10^{-34} \text{ J s}$
$\hbar = h/2\pi$	$1.05457 \times 10^{-34} \text{ J s}$
c	$2.99793 \times 10^8 \text{ m/s}$
e	$1.60217 \times 10^{-19} \text{ C}$
m_e	$9.10938 \times 10^{-31} \text{ kg}$
ϵ_0	$8.85419 \times 10^{-12} \text{ C}^2 \text{J}^{-1} \text{m}^{-1}$
R (Rydberg const.)	109737.3 cm^{-1}

- 1.) (17 points) When light with a wavelength of 3000 \AA shines on a metal surface, the kinetic energy of ejected electrons is $2.960 \times 10^{-19} \text{ J}$. When light with a wavelength of 5400 \AA shines on the same metal surface, the kinetic energy of ejected electrons is $1.362 \times 10^{-21} \text{ J}$. From this information, determine the value of Planck's constant (in units of J s). [$1 \text{ \AA} = 10^{-10} \text{ m}$]

For the photoelectric effect, $KE = h\nu - \phi$. From this equation, we can see that a plot with kinetic energy on the y -axis and frequency on the x -axis yields a straight line with the slope equal to Planck's constant h . Therefore, one way to determine Planck's constant from the data given is to calculate the slope,

$$h = \text{slope} = \frac{\Delta KE}{\Delta \nu}$$

The kinetic energy values are given in the problem, but the frequencies must be calculated from the wavelengths.

$$\begin{aligned} \nu_1 &= \frac{c}{\lambda_1} \\ &= \frac{2.99793 \times 10^8 \text{ m/s}}{(3000 \text{ \AA}) \left(\frac{10^{-10} \text{ m}}{1 \text{ \AA}} \right)} \\ \nu_1 &= 9.993 \times 10^{14} \text{ s}^{-1} \end{aligned}$$

$$\begin{aligned} \nu_2 &= \frac{c}{\lambda_2} \\ &= \frac{2.99793 \times 10^8 \text{ m/s}}{(5400 \text{ \AA}) \left(\frac{10^{-10} \text{ m}}{1 \text{ \AA}} \right)} \\ \nu_2 &= 5.552 \times 10^{14} \text{ s}^{-1} \end{aligned}$$

Physical Chemical Self Test Solution

JL Elias

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