

**C11- PHYSICAL SCIENCE**  
Third Monthly Test Reviewer

Coverage: **MODULE 1**  
**UNIT 1: Atomic Theory**  
**UNIT 2: The Discovery and Synthesis of Chemical Elements**

**UNIT 1: Atomic Theory**

- Atoms are composed of **protons, neutrons, and electrons.**



Table 2.3 Properties of Subatomic Particles

Particle	Symbol	Mass (kg)	Relative Mass (proton =1)	Relative Charge
Proton	$p^+$	$1.675 \times 10^{-27}$	1	+1
Neutron	$n^0$	$1.675 \times 10^{-27}$	1	0
Electron	$e^-$	$9.109 \times 10^{-31}$	0.00055	-1

- An atom is mostly empty space with a highly dense, **positive** center called **nucleus**.
- For a neutral atom, the **number of protons is equivalent to the number of electrons.**  
(Lithium has 3 protons therefore it also has 3 electrons)
- The **atomic number** of an element is also equivalent to its number of **protons.**  
(Nitrogen has an atomic number of 7 therefore it has 7 protons in its nucleus)
- The **atomic mass** (mass number) is primarily determined by the number of protons and **neutrons.**

- Isotopes** are atoms with the **same atomic number** but different mass numbers. They **differ in the number of neutrons.**  
(Carbon-12, Carbon-13, Carbon-14)
- Ion** is an element or atom that has a **positive or negative net charge.** It is charged because the number of electrons does not equal the number of protons in the atom.

Two types:

- Cation**- positive ion (loses electron)
- Anion**- negative ion (gains electron)



- Take note that when an **atom is not electrically neutral**, it means that it does not have the same number of protons and electrons.
- When an atom or element is **positively charged**, it means that it has **more protons** in its nucleus.

**Historical Development of Atom**

- Democritus & Leucippus:**  
**Atomism:** All matter is made up of very small, indivisible particles called atoms (atomos).
- Aristotle:** Four elements in nature (air, fire, water, and earth)
- John Dalton: Billiard Ball Model/Solid Sphere Model:**  
**Atomic Theory:** All matter is made up of very small, indivisible particles called atoms.

# Phisical Science Gradeexemplar Caps 2014

**Christian Drosten**



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