

The Periodic Law

SECTION 2

Select the correct answer for the following questions for the space provided.

1. _____ for the modern periodic table, elements are ordered
 - (a) according to increasing atomic mass
 - (b) according to Mendeleev's original design
 - (c) according to increasing atomic number
 - (d) based on how they were discovered
2. _____ Mendeleev's original short periodic table gave better chemical properties of elements appeared in regular intervals and whereas elements were arranged according to increasing
 - (a) density
 - (b) atomic number
 - (c) atomic mass
 - (d) electronegativity
3. _____ The modern periodic law states that
 - (a) all new elements with the same properties are listed in the same group
 - (b) the physical and chemical properties of elements are functions of atomic number
 - (c) elements exhibit properties of both periods and groups
 - (d) the chemical properties of elements can be grouped according to periodicity but physical properties cannot
4. _____ The discovery of the noble gases changed Mendeleev's periodic table
 - (a) adding a new
 - (b) period
 - (c) series
 - (d) group
5. _____ The noble gases that property of the noble gases is that they are
 - (a) unstable
 - (b) inert
 - (c) radioactive
 - (d) highly reactive
6. _____ Likewise, the first element in Group 1 has an atomic number of 3. The second element in this group has an atomic number of
 - (a) 4
 - (b) 11
 - (c) 19
 - (d) 37
7. How groups in Mendeleev's periodic table of 18 and the atomic number of 55
 - (a) 11. How many periods are in this series?
 - (b) 19. How many elements are in this group?
 - (c) 19. How many elements are in this group?
 - (d) 11. How many elements are in this group?

Periodic Law Chapter 5 Test Answers

Pieter Thyssen, Arnout Ceulemans

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author names.

Periodic Law Chapter 5 Test Answers:

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trail follows the lines of parentage linking groups upstream to the unitary symmetry of the eightfold way of quarks and to the four dimensional symmetry of the hydrogen atom Along the way the theory of symmetry groups is gradually introduced with special emphasis on graphical representations The final challenge is to open up the structure of Mendeleev s table which goes beyond the symmetry of the hydrogen atom Breaking this symmetry to accommodate the multi electron atoms requires to leave the common ground of linear algebras and explore the potential of non linearity *Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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