

Group name*		
Hydrogen & the alkali metals	14	Crystallogens
Alkaline earth metals	15	Psithogens
Coinage metals	16	Chalcogens
Volatile metals	17	Halogens
Isotopans	18	Helium atom

*Groups 1 to 20 are named after their first members i.e. Group 1 is the 'Island Group'

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☐ **Radioactive**
 Element is
 likely to be either
 a metalloid or a
 good transition
 metal

Metals
Metallic looking, brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionization energies (700–1,000 kJ/mol), and electronegativity values (1.8–2.2). Usually behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

Intermediate nonmetals
H and N are colorless gases; C, P*, and Se are metallic looking solids while Si is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionization energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidizing agents*, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).

* as black P, the most stable form in ambient conditions

Group 1
This starts as Be²⁺, and then bifurcates into a La branch (i.e., La–Ac), and a f block (i.e., f–L). Thanks going down Be–F, and the:

- La branch, resembles those of groups 1–2
- f block, more resembles those of groups 4–10.

Corrosive nonmetals
O, F, and Cl are, respectively, colorless, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have a high ionization energies, electron affinities, and electronegativity values, and be relatively strong oxidizing agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acids.

Noble gases
Colorless, odorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePtF_6 , was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesized in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Periodic Table Guide Answers

Lauren Gardner



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