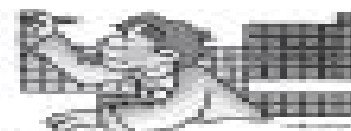


ACTIVITY: Periodic Trends – WEBQUEST!

This activity will provide an opportunity to examine the periodic table more closely. Use this interactive periodic table (see link below) to explore trends of atomic radius, ionization energy, and electronegativity across a period and down a group on the periodic table.



Go to <http://www.ck12.com> and click on the properties tab near the upper left of the screen.

Part 1 – Atomic Radius:

(click on radius / calculated radius in the options in the middle of the screen)

1) Define atomic radius: look up in your notes or text or online)



2) Examine the trend in atomic radius from left to right across a period by clicking on all the elements in the 2nd period. List their atomic radii below (pm = picometer or 1×10^{-12} m).

Li _____, Be _____, B _____, C _____, N _____, O _____, F _____, Ne _____

3) What is the trend in atomic radius from left to right across a period?

4) Rationalize / explain this trend with respect to the nuclear attraction to the electron cloud. (look up in your notes or text or online if you need to)

5) Examine the trend in atomic radius from top to bottom down a group by clicking on all the elements in the 1st group. List their atomic radii below (pm = picometer or 1×10^{-12} m).

Li _____, Na _____, K _____, Rb _____, Cs _____, Fr _____

6) What is the trend in atomic radius from top to bottom down a group?

7) Rationalize / explain this trend with respect to the nuclear attraction to the electron cloud. (look up in your notes or text or online if you need to)



8) Without looking at their atomic radii on the periodic table, try to organize the following elements in order of increasing atomic radius (Ca, Hc, Sr, Cl, Ba). Then check your answer using the table!

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