

- 2.77 (a) Li^{+} , alkali metals always have a +1 charge in ionic compounds.
 (b) S^{2-} .
 (c) I^{-} , halogens have a -1 charge in ionic compounds.
 (d) N^{3-} .
 (e) Al^{3+} , aluminum always has a +3 charge in ionic compounds.
 (f) Ca^{2+} , alkali metals always have a +1 charge in ionic compounds.
 (g) Mg^{2+} , alkaline earth metals always have a +2 charge in ionic compounds.

2.78 The symbol ^{23}Na provides more information than Na . The mass number plus the chemical symbol identifies a specific isotope of Na (sodium) while combining the atomic number with the chemical symbol tells you nothing new. Can other isotopes of sodium have different atomic numbers?

2.79 The binary Group 7A element acids are: HF, hydrofluoric acid; HCl, hydrochloric acid; HBr, hydrobromic acid; HI, hydroiodic acid. Oxacids containing Group 7A elements (using the specific examples for chlorine) are: HClO_4 , perchloric acid; HClO_3 , chloric acid; HClO_2 , chlorous acid; HClO , hypochlorous acid.

Examples of oxacids containing other Group A-block elements are: H_2BO_3 , boric acid (Group 3A); H_2CO_3 , carbonic acid (Group 4A); HNO_3 , nitric acid and H_3PO_4 , phosphoric acid (Group 5A); and H_2SO_4 , sulfuric acid (Group 6A). Hydrofluoric acid, HF , is an example of a binary Group 6A acid while HCN , hydrocyanic acid, contains both a Group 4A and 5A element.

2.80 Mercury (Hg) and bromine (Br)

2.81 (a)	Isotope	^{214}Po	^{210}Po	^{210}At	^{210}Bi	^{210}Po
	No. Protons	84	84	85	83	84
	No. Neutrons	130	126	125	127	126
(b)	neutron/proton ratio	1.55	1.50	1.47	1.53	1.51

The neutron/proton ratio increases with increasing atomic number.

2.82 H_2 , N_2 , O_2 , F_2 , Cl_2 , Br_2 , I_2 , Ne , Ar , Kr , Xe , Rn

2.83 Cu, Ag, and Au are fairly chemically unreactive. This makes them especially suitable for making coins and jewelry, that you want to last a very long time.

2.84 They do not have a strong tendency to form compounds. Helium, neon, and argon are chemically inert.

2.85 Magnesium and calcium are also alkaline earth metals. You should expect the charge of the metal to be the same (+2). MgO and CaO .

2.86 All isotopes of radium are radioactive. It is a radioactive decay product of uranium-238. Radium itself does not occur naturally on Earth.

- 2.87 (a) Berkelium (Berkley, CA); Europium (Europe); Francium (France); Scandium (Scandinavia);
 Ytterbium (Ytterby, Sweden); Yttrium (Ytterby, Sweden).
 (b) Einsteinium (Albert Einstein); Fermium (Enrico Fermi); Curium (Marie and Pierre Curie);
 Mendelevium (Dmitri Mendeleev); Lawrencium (Ernest Lawrence).
 (c) Arsenic, Cesium, Chlorine, Chromium, Iodine.

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Michael Rajnik



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