Name:	Date:

## Calculating Average Rates of Reaction

You are required to show work for all calculations.

During the combustion of methane, CH<sub>4</sub>, shown by the reaction

$$CH_4(g) + O_2(g) \rightarrow CO_2(g) + H_2O(g)$$

The concentration of methane was measured at various time intervals and the following results were obtained:

Time	[CH <sub>4</sub> ]
(s)	(mol - l <sup>-1</sup> )
10	2.40
20	1.20
30	0.80
40	0.60

Calculate the average rate of loss of methane during the 10 to 40 second time period.

#### Solution:

$$rate = \frac{\Delta[CH_4]}{\Delta t}$$

$$= \frac{[0.60 - 2.4]mol}{(40 - 10)s \times L}$$

$$= \frac{-1.8mol}{30s \times L}$$

$$= -0.06 \frac{mol}{L \times s}$$

Note: the answer is negative because we are calculating the rate of loss of methane.

# Calculating Reaction Rates using Stoichiometry

## Example:

In the following decomposition reaction,

$$2 N_2O_5 \rightarrow 4 NO_2 + O_2$$

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**Lingsheng Yao** 

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