

Reaction Rates and Chemical Equilibrium



Glossary

TERM	DEFINITION
activation energy	the minimum amount of energy needed to start a chemical reaction
catalyst	a substance that lowers the activation energy of a reaction but is otherwise not part of the reaction
chemical equation	the symbolic representation of a chemical reaction can be written as
chemical equilibrium	a state of a reversible reaction in which the forward and reverse reactions occur at the same rate and the quantities of reactants and products do not change
chemical reaction	a process that involves the rearrangement of atoms, molecules, or both to form new substances
concentration	the amount of solute in a given amount of solvent or solution
endothermic reaction	a chemical reaction that absorbs energy
energy	the property of a body or a system that enables it to do work or cause a change
exothermic reaction	a chemical reaction that releases energy
product	a substance formed in a chemical reaction
reactant	a substance that is present at the beginning of a chemical reaction and is transformed by the reaction
reaction rate	the rate at which reactants are used up or products are formed in a chemical reaction
reversible reaction	a chemical reaction that occurs simultaneously in the forward and reverse directions, converting reactants to products and products to reactants

Objective

In this lesson, you will predict how altering the environment in which a chemical reaction takes place changes the rate of the reaction and the amounts of products.

Reaction Rates

In a chemical reaction, atoms rearrange themselves by breaking and forming new bonds. The atoms of the reactants rearrange to form the products, which are made of the same atoms as the reactants. The time taken for reactants to convert to products is the reaction rate. As a reaction progresses, the amount of reactants decreases and the amount of products increases.

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chemistry Additionally all scientists can benefit from improved critical thinking and problem solving skills that are developed from the study of organic chemistry Organic chemistry like any skill is best learned by doing It is difficult to learn by rote memorization and true understanding comes only from concentrated reading and working as many problems as possible In fact problem sets are the best way to ensure that concepts are not only well understood but can also be applied to real world problems in the work place Helps readers learn to categorize analyze and solve organic chemistry problems at all levels of difficulty Hundreds of fully worked practice problems all with solutions Key concept summaries for every chapter reinforces core content from the companion book

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