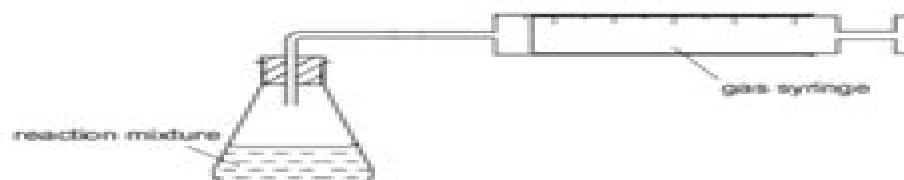


Reaction rates

0620 Cambridge IGCSE Chemistry Specimen Papers (2016)

Paper 1 Question 15

15. The apparatus shown can be used to measure the rate of some chemical reactions.



For which two reactions would this apparatus be suitable?

- | | |
|------------|------------------------------------------------------------------------------------------------------------------------------------------------|
| reaction 1 | $\text{AgNO}_3(\text{aq}) + \text{HCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{HNO}_3(\text{aq})$ |
| reaction 2 | $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ |
| reaction 3 | $\text{MgO}(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{MgCl}_2(\text{aq}) + \text{H}_2\text{O}(\text{l})$ |
| reaction 4 | $\text{ZnCO}_3(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{ZnCl}_2(\text{aq}) + \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$ |

- A 1 and 2 B 1 and 3 C 2 and 4 D 3 and 4

Paper 1 Question 16

16. A student investigates the rate of reaction between magnesium and excess sulfuric acid.

The volume of hydrogen given off in the reaction is measured over time.

The graph shows the results of two experiments, R and S.



Which change in conditions would cause the difference between R and S?

- A A catalyst is added in S.
 B The acid is more concentrated in R than in S.
 C The magnesium is less finely powdered in R than in S.
 D The temperature in R is lower than in S.

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Sebastian Brünink



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