Rate Laws and Reaction Order

Rate Law
is a relation between the rate of reaction to the concentrations of the reactants

In general:

the rate law equation is:

 the relationship among rate, the rate constant, the initial concentration of reactants, and the orders of reaction with respect to the reactants

Rate =
$$k[A]^y[B]^y[C]^z$$
...

where k is the rate constant, i.e., it is a proportionality constant.

where sis the order of the reaction with respect to A.

s the order of the reaction with respect to B.

zis the order of the reaction with respect to C.

 the exponent used to describe the relationship between the initial concentration of a particular reactant and the aste of the reaction

The sum of x + y + z + ... is called the overall order.

Reaction Order And Rate Laws Lab Answers

RS Peters

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