


# AQA A Level Chemistry Presentations

## LECTURE 36 - RATES OF REACTION



The diagram shows a laboratory setup for measuring the rate of a chemical reaction. A conical flask containing a blue liquid (hydrochloric acid) and small pieces of a grey solid (magnesium) is sealed with a stopper. A delivery tube is inserted through the stopper, leading from the flask to an inverted gas syringe. Bubbles of gas are shown entering the syringe from the tube. Labels with leader lines identify the 'pieces of Mg' at the bottom of the flask, the 'acid solution' (the blue liquid), the 'bubbles of H<sub>2</sub> gas under pressure' rising from the magnesium, and the 'H<sub>2</sub> gas collected in gas syringe' at the top of the inverted syringe.

The reaction between HCl and Mg produces H<sub>2</sub> gas

$$\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$$

The volume of hydrogen gas produced is recorded at regular intervals. The rate of reaction can be determined by how fast the gas volume increases.

# Reaction Rates Lesson 36 Answers

**LL Leslie**



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