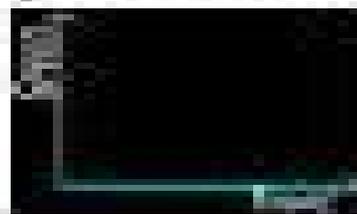


2<sup>nd</sup> Year Honours BSc in Applied Science, University of Exeter, UK  
 CH101

This assignment asks you to investigate the kinetics of the reaction between  $\text{H}_2\text{O}_2$  and  $\text{I}^-$ . This reaction is very fast and so you will need to use a technique that allows you to measure the concentration of  $\text{I}_2$  as it is formed. You will be asked to determine the order of reaction with respect to  $\text{H}_2\text{O}_2$  and  $\text{I}^-$ .

Introduction

- 1) What is the rate of reaction? **1/2** Rate of reaction  
 Change in concentration of a reactant or product over time
- 2) What is the difference between average and instantaneous rate? **1/2**  
 Average rate is the change in concentration over a period of time.  
 Instantaneous rate is the rate at a particular point in time.
- 3) What is the difference between average and instantaneous rate? **1/2**  
 Average rate is the change in concentration over a period of time.  
 Instantaneous rate is the rate at a particular point in time.
- 4) How do you determine the order of reaction with respect to a reactant? **1/2**  
 Change in concentration of a reactant over time.  
 Change in concentration of a reactant over time.  
 Change in concentration of a reactant over time.
- 5) What are the different ways of determining the order of reaction? **1/2**  
 Change in concentration of a reactant over time.  
 Change in concentration of a reactant over time.  
 Change in concentration of a reactant over time.
- 6) How does the rate of reaction change with concentration? **1/2**  
 Change in concentration of a reactant over time.
- 7) How does the rate of reaction change with temperature? **1/2**  
 Change in concentration of a reactant over time.
- 8) What is the effect of a catalyst on the rate of reaction? **1/2**  
 Change in concentration of a reactant over time.



# Second Semester Icp Exam

**Christian G. Meyer**



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