

11

CHEMICAL REACTIONS

SECTION 11.1 DESCRIBING CHEMICAL REACTIONS (pages 321–329)

This section explains how to write equations describing chemical reactions using appropriate symbols. It also describes how to write balanced chemical equations when given the names or formulas of the reactants and products in a chemical reaction.

► Writing Chemical Equations (pages 321–323)

1. A chemical reaction occurs when one or more _____ change into one or more new substances called _____.
2. The arrow in a reaction means _____.
3. Is the following sentence true or false? When there are two or more reactants or products, they are separated by an arrow. _____
4. Write a word equation that describes the following reactions.
 - a. Acetylene reacts with oxygen to produce carbon dioxide and water.

 - b. When heated, mercury(II) oxide reacts to form mercury and oxygen.

5. What is a chemical equation?

6. A chemical reaction that shows only the formulas, but not the relative amounts of the reactants and products is a(n) _____.
7. Identify the reactant(s) and product(s) in the chemical equation $\text{Li} + \text{Br}_2 \longrightarrow \text{LiBr}$.
 - a. reactant(s) _____
 - b. product(s) _____
8. Circle the letter of each statement that is true about a catalyst.
 - a. A catalyst is the new material produced as a result of a chemical reaction.
 - b. A catalyst is not used up in a chemical reaction.
 - c. A catalyst adds heat to a chemical reaction.
 - d. A catalyst speeds up a chemical reaction.

Section 11 1 Describing Chemical Reactions Answers Key

**Dean R. Hess, Neil R. MacIntyre, Shelley
C. Mishoe, William F. Galvin**

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Science Explorer Physical Science Michael J. Padilla, Ioannis Miaculis, Martha Cyr, *Essential Biochemistry* Charlotte W. Pratt, Kathleen Cornely, 2023-08-25 *Essential Biochemistry* 5th Edition is comprised of biology pre med and allied health topics and presents a broad but not overwhelming base of biochemical coverage that focuses on the chemistry behind the biology This revised edition relates the chemical concepts that scaffold the biology of biochemistry providing practical knowledge as well as many problem solving opportunities to hone skills Key Concepts and Concept Review features help students to identify and review important takeaways in each section

Chemical Reactivity in Confined Systems Pratim Kumar Chattaraj, Debduitta Chakraborty, 2021-08-23 An insightful analysis of confined chemical systems for theoretical and experimental scientists *Chemical Reactivity in Confined Systems Theory and Applications* presents a theoretical basis for the molecular phenomena observed in confined spaces The book highlights state of the art theoretical and computational approaches with a focus on obtaining physically relevant clarification of the subject to enable the reader to build an appreciation of underlying chemical principles The book includes real world examples of confined systems that highlight how the reactivity of atoms and molecules change upon encapsulation Chapters include discussions on recent developments related to several host guest systems including cucurbit n uril ExBox 4 clathrate hydrates octa acid cavitand metal organic frameworks MOFs covalent organic frameworks COFs zeolites fullerenes and carbon nanotubes Readers will learn how to carry out new calculations to understand the physicochemical behavior of confined quantum systems Topics covered include A thorough introduction to global reactivity descriptors including electronegativity hardness and electrophilicity An exploration of the Fukui function as well as dual descriptors higher order derivatives and reactivity through information theory A practical discussion of spin dependent reactivity and temperature dependent reactivity Concise treatments of population analysis reaction force electron localization functions and the solvent effect on reactivity Perfect for academic researchers and graduate students in theoretical and computational chemistry and confined chemical systems *Chemical Reactivity in Confined Systems Theory and Applications* will also earn a place in the libraries of professionals working in the areas of catalysis supramolecular chemistry and porous materials

Introductory Chemistry Charles H. Corwin, 2005 For one semester courses in Basic Chemistry Introduction to Chemistry and Preparatory Chemistry and the first term of Allied Health Chemistry This text is carefully crafted to help students learn chemical skills and concepts more effectively Corwin covers math and problem solving early in the text he builds student confidence and skills through innovative problem solving pedagogy and technology formulated to meet student needs

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