

2.3

Carbon Compounds

THINK ABOUT IT In the early 1800s, many chemists called the compounds created by organisms “organic,” believing they were fundamentally different from compounds in nonliving things. Today we understand that the principles governing the chemistry of living and nonliving things are the same, but the term “organic chemistry” is still around. Today, organic chemistry means the study of compounds that contain bonds between carbon atoms, while inorganic chemistry is the study of all other compounds.

The Chemistry of Carbon

What elements does carbon bond with to make up life's molecules?

Why is carbon so interesting that a whole branch of chemistry should be set aside just to study carbon compounds? There are two reasons for this. First, carbon atoms have four valence electrons, allowing them to form strong covalent bonds with many other elements.

Carbon can bond with many elements, including hydrogen, oxygen, phosphorus, sulfur, and nitrogen to form the molecules of life. Living organisms are made up of molecules that consist of carbon and these other elements.

Even more important, one carbon atom can bond to another, which gives carbon the ability to form chains that are almost unlimited in length. These carbon-carbon bonds can be single, double, or triple covalent bonds. Chains of carbon atoms can even close up on themselves to form rings, as shown in Figure 2-12. Carbon has the ability to form millions of different large and complex structures. No other element even comes close to matching carbon's versatility.

FIGURE 2-12 Carbon Structures Carbon can form single, double, or triple bonds with other carbon atoms. Each line between atoms in a molecular drawing represents one covalent bond. **Observing** How many covalent bonds are there between the two carbon atoms in acetylene?



Methane

Acetylene

Butadiene

Benzene

Isooctane

Key Questions

What elements does carbon bond with to make up life's molecules?

What are the functions of each of the four groups of macromolecules?

Vocabulary

monomer • polymer •
carbohydrate •
monosaccharide •
lipid • nucleic acid •
nucleotide • protein •
amino acid

Taking Notes

Compare/Contrast Table As you read, make a table that compares and contrasts the four groups of organic compounds.

Getting Started

Objectives

2.3.1 Describe the unique qualities of carbon.

2.3.2 Describe the structures and functions of each of the four groups of macromolecules.

Student Resources

Study Workbooks A and B, 2.3 Worksheets

Spanish Study Workbook, 2.3 Worksheets

Lab Manual B, 2.3 Data Analysis Worksheet

Lesson Overview • Lesson Notes
• Assessment: Self-Test, Lesson Assessment

For corresponding lesson in the Foundation Edition, see pages 37–41.

Activate Prior Knowledge

Ask students to name compounds they know that contain carbon. As they make suggestions, create a class list on the board. Use their list to start a discussion on the wide variety of carbon compounds that exist in nature.

Answers

FIGURE 2-12 three



NATIONAL SCIENCE EDUCATION STANDARDS

UNIFYING CONCEPTS AND PROCESSES

L, V

CONTENT

B.1, B.2, C.1.a, C.1.c, C.1.d, C.2.a, C.5.b

Ubd Teach for Understanding

ENDURING UNDERSTANDING The process of science helps biologists investigate how nature works at all levels, from the molecules in cells to the biosphere.

GUIDING QUESTION How do organisms use different types of carbon compounds?

EVIDENCE OF UNDERSTANDING After completing the lesson, give students the following assessment to show they understand the functions of the different types of carbon compounds in organisms. Ask each student to create a four-page brochure, with each page devoted to one of the four major types of carbon compounds in living things. For each type of compound, students should diagram its general structure and describe its functions in organisms.

Section 2 3 Carbon Compounds Answer Key

EduGorilla Prep Experts

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Manual of Chemistry: Chemistry of carbon-compounds; or, Organic chemistry Henry Watts, 1886

The Pyrolysis of Carbon Compounds Charles Dewitt Hurd, 1929

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