

Name: _____

SBI 4U

Remember this stuff?
Review: Atoms and Bonding
Refer to text pages 6-12

Atom – the fundamental unit of all matter. It is the smallest particle of matter that has all the properties of a particular element.

- The atom is composed of 3 subatomic particles:
 - Protons: positive charge, found in nucleus
 - Electrons: negative charge, outside nucleus
 - Neutrons: no charge, found in nucleus
- An atom contains no electric charge:
 - The negative charge of one electron is as strong as the positive charge of one proton
 - The number of protons and electrons in an atom is equal, so the charges cancel each other out.

Molecule – the combination of two or more atoms (of the same or different elements)

Examples: Cl_2 , H_2O , NaCl

Element – substances containing only one type of atom

Examples: Al , Au , Na

Compound – a pure substance that cannot be broken down into simpler substances

- contains two or more different atoms in specific proportions

Examples: HCl , NaOH

Atomic Number – represents the number of protons in an atom

Atomic Mass – the total number of protons and neutrons in an atom

Section 2 Naming Molecules Science Notebook Answers

Dr Chandresh Agrawal



Section 2 Naming Molecules Science Notebook Answers:

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 - Chapter 2: Essential Elements of Section 2 Naming Molecules Science Notebook Answers
 - Chapter 3: Section 2 Naming Molecules Science Notebook Answers in Everyday Life
 - Chapter 4: Section 2 Naming Molecules Science Notebook Answers in Specific Contexts
 - Chapter 5: Conclusion
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