

# KEY

## Percent, Actual, and Theoretical Yield



a) I began this reaction with 20 grams of lithium hydroxide. What is my theoretical yield of lithium chloride?

$$\frac{20 \text{ g LiOH}}{23.95 \text{ g LiOH}} \times \frac{1 \text{ mol LiOH}}{1 \text{ mol LiOH}} \times \frac{42.39 \text{ g LiCl}}{1 \text{ mol LiCl}} = \boxed{35.40 \text{ g LiCl}}$$

b) I actually produced 6 grams of lithium chloride. What is my percent yield?

$$\frac{6 \text{ g LiCl}}{35.40 \text{ g LiCl}} \times 100 = \boxed{16.9\%}$$



a) If I start with 5 grams of  $\text{C}_3\text{H}_8$ , what is my theoretical yield of water?

$$\frac{5 \text{ g C}_3\text{H}_8}{44.1 \text{ g C}_3\text{H}_8} \times \frac{1 \text{ mol C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{8.175 \text{ H}_2\text{O}}$$

b) I got a percent yield of 75%. How many grams of water did I get?

$$(8.175 \text{ H}_2\text{O})(.75) = \boxed{6.13 \text{ g H}_2\text{O}}$$



My theoretical yield of barium chloride was 10.7 grams. If my actual yield was 9.5 grams, what was my percent yield?

$$\frac{9.5}{10.7} \times 100 = \boxed{88\%}$$



What is my theoretical yield of sodium oxide if I start with 20 grams of calcium oxide?

$$\frac{20 \text{ g CaO}}{56.08 \text{ g CaO}} \times \frac{1 \text{ mol CaO}}{1 \text{ mol CaO}} \times \frac{61.98 \text{ g Na}_2\text{O}}{1 \text{ mol Na}_2\text{O}} = \boxed{22.10 \text{ g Na}_2\text{O}}$$

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