

Complete these problems on a separate sheet of paper.

1. Determine the percent Oxygen in Potassium Perchlorate.



Molar Mass is 138.547 g  $\text{KClO}_4$

4 moles of O in  $\text{KClO}_4$

Molar mass of O is 15.999

4 moles O = 63.996 g

$$\text{Percent Composition of O} = \frac{63.996 \text{ g}}{138.547 \text{ g}} \times 100 = 46.19 \%$$

2. Determine the percent Aluminum in Aluminum Sulfate.



Molar Mass is 342.15 g

2 moles of Al in  $\text{Al}_2(\text{SO}_4)_3$

Molar Mass of Al = 26.982 g

2 moles Al = 53.964 g

$$\text{Percent Composition of Al} = \frac{53.964 \text{ g}}{342.15 \text{ g}} \times 100 = 15.77 \%$$

3. Determine the percent Iron in Iron II Phosphite.



Molar Mass is 325.741 g

2 moles of Fe in  $\text{Fe}_2(\text{PO}_3)_2$

Molar Mass of Fe = 55.933 g

2 moles of Fe = 111.866 g

$$\text{Percent Composition of Fe} = \frac{111.866 \text{ g}}{325.741 \text{ g}} \times 100 = 34.34 \%$$

4. A 4.829 gram sample of Magnesium Sulfate Septahydrate (Epsom Salt). If the sample is heated and all of the water is driven off, how much anhydrous salt will remain?



$$\frac{4.829 \text{ g MgSO}_4 \cdot 7\text{H}_2\text{O}}{1} \times \frac{1 \text{ mole MgSO}_4 \cdot 7\text{H}_2\text{O}}{246.472 \text{ g MgSO}_4 \cdot 7\text{H}_2\text{O}} \times \frac{1 \text{ mole MgSO}_4}{1 \text{ mole MgSO}_4 \cdot 7\text{H}_2\text{O}} \times \frac{120.367 \text{ g MgSO}_4}{1 \text{ mole MgSO}_4} = 2.358 \text{ g MgSO}_4$$

5. A 7.028 gram sample of a sodium sulfate hydrate is heated. After driving off all the water, 3.100 grams of the anhydrous salt is left. What is the formula of the hydrate?



$$\text{Moles of Anhydrous Salt} = \frac{3.100 \text{ g Na}_2\text{SO}_4}{1} \times \frac{1 \text{ mole Na}_2\text{SO}_4}{142.042 \text{ g Na}_2\text{SO}_4} = 0.02182 \text{ moles Na}_2\text{SO}_4$$

# Percent Composition Answers

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Velavancorresponding**



## **Percent Composition Answers:**

## Unveiling the Power of Verbal Artistry: An Mental Sojourn through **Percent Composition Answers**

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