

# Percent Yield Practice

3 pages + 3-page Answer KEY + Periodic Table

Chemistry

## PERCENT YIELD

Show all work and circle or highlight your final answer.

$$\text{percent yield (\%)} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

- 1 16.8 grams of sodium reacts with excess chlorine gas.



The theoretical yield of sodium chloride is 25.5 grams. The actual yield was 23.5 grams. Calculate the percent yield of sodium chloride.

Name \_\_\_\_\_

Date \_\_\_\_\_

- 2 Hydrogen gas ( $\text{H}_2$ ) reacts with oxygen gas ( $\text{O}_2$ ) to produce water ( $\text{H}_2\text{O}$ ). 5.8 grams of hydrogen reacts with excess oxygen gas.



What is the theoretical yield?

- 3 8.8 grams of carbon reacts with excess oxygen gas.



The theoretical yield of carbon dioxide is 25.5 g. Calculate the percent yield of carbon dioxide.

- 4 1.8 grams of zinc reacts with excess sulfur.



The theoretical yield of zinc sulfide is 3.5 grams. Calculate the percent yield of zinc sulfide.

- 5 The actual yield was 28.8 grams. Calculate the percent yield.

- 6 Magnesium ( $\text{Mg}$ ) reacts with hydrochloric acid ( $\text{HCl}$ ) to produce magnesium chloride ( $\text{MgCl}_2$ ) and hydrogen gas ( $\text{H}_2$ ). 10.0 grams of magnesium reacts with excess hydrochloric acid.



What is the theoretical yield?

- 7 1.0 grams of phosphorus reacts with excess chlorine gas.



The theoretical yield of phosphorus trichloride is 3.5 grams. Calculate the percent yield of phosphorus trichloride.

- 8 The actual yield was 40.2 grams. Calculate the percent yield.

Name \_\_\_\_\_

Date \_\_\_\_\_

- 9 Calcium carbonate ( $\text{CaCO}_3$ ) reacts with hydrochloric acid ( $\text{HCl}$ ) to produce calcium chloride ( $\text{CaCl}_2$ ), carbon dioxide ( $\text{CO}_2$ ), and water ( $\text{H}_2\text{O}$ ). 15.0 grams of calcium carbonate reacts with excess hydrochloric acid.



What is the theoretical yield of calcium chloride?

$$\begin{array}{r} 15.0 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} \times \frac{146.97 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 13.9 \text{ g CaCl}_2 \end{array}$$

$$\begin{array}{r} \text{Molar Mass CaCl}_2 \\ \text{Ca} = 40.08 \\ \text{Cl} = 35.45 \\ 2 \times 35.45 = 70.90 \\ 40.08 + 70.90 = 110.98 \text{ g} \end{array}$$

10 The actual yield was 13.9 grams. Calculate the percent yield of calcium chloride.

$$\begin{array}{r} 13.9 \text{ g CaCl}_2 \times \frac{1 \text{ mol CaCl}_2}{146.97 \text{ g CaCl}_2} \times \frac{1 \text{ mol CaCO}_3}{1 \text{ mol CaCl}_2} \times \frac{100.09 \text{ g CaCO}_3}{1 \text{ mol CaCO}_3} = 9.5 \text{ g CaCO}_3 \end{array}$$

$$\begin{array}{r} \text{Molar Mass CaCO}_3 \\ \text{Ca} = 40.08 \\ \text{C} = 12.01 \\ 3 \times 16.00 = 48.00 \\ 40.08 + 12.01 + 48.00 = 100.09 \text{ g} \end{array}$$

- 11 The actual yield was 25.5 grams. Calculate the percent yield of calcium chloride.

$$\text{percent yield (\%)} = \frac{25.5 \text{ g}}{100.09 \text{ g}} \times 100 = 25.5\%$$

- 12 Potassium ( $\text{K}$ ) reacts with iodine ( $\text{I}_2$ ) to produce potassium iodide ( $\text{KI}$ ). 10.0 grams of potassium reacts with excess iodine.



What is the theoretical yield?

$$\begin{array}{r} 10.0 \text{ g K} \times \frac{1 \text{ mol K}}{39.10 \text{ g K}} \times \frac{2 \text{ mol KI}}{2 \text{ mol K}} \times \frac{166.10 \text{ g KI}}{1 \text{ mol KI}} = 166.10 \text{ g KI} \end{array}$$

$$\begin{array}{r} \text{Molar Mass KI} \\ \text{K} = 39.10 \\ \text{I} = 126.90 \\ 39.10 + 126.90 = 166.00 \text{ g} \end{array}$$

YIELD

ANSWER KEY INCLUDED

# Percent Yield Example Answer Key

**Arshad Iqbal**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## **Percent Yield Example Answer Key:**

Gcse Biology Stugy Guide Daniel W Foster Professor of Medical Ethics John Sadler,2007-10 Written by examiners and practising teachers each book in this series contains activities and useful features intended to aid understanding Knowledge is tested throughout with progress checks at the end of every chapter and practice questions at the end of each section

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Steven M. Rice, 2018-12-11 Get ready to qualify for the security industry job of your dreams The Securities Industry Essentials SIE Exam is a new test administered by FINRA beginning in October 2018 This exam is required as a prerequisite to each series level knowledge exam such as Series 7 If you're prepping for the exam you need a trusted resource to ensure your very best performance Securities Industry Essentials Exam For Dummies with Online Practice gives you everything you need to score high on this important exam With two practice tests in the book plus two bonus tests online you can practice your way to a calm and confident experience on exam day Take 4 full length practice tests with answers and full explanations Get 1 year access to practice and tests online Find strategies and tips for breaking into the securities industry Increase your chances of scoring higher SIE scores are expected to be a critical factor in determining qualification to enter the securities industry so the stakes are high With the help of this book you'll up your chances of breaking into this field and landing your dream job

**Congressional Record** United States. Congress, 2000 The Congressional Record is the official record of the proceedings and debates of the United States Congress It is published daily when Congress is in session The Congressional Record began publication in 1873 Debates for sessions prior to 1873 are recorded in The Debates and Proceedings in the Congress of the United States 1789 1824 the Register of Debates in Congress 1824 1837 and the Congressional Globe 1833 1873

**Climate-resilience policies and investments for Egypt's agriculture sector: Sustaining productivity and food security** Perez, Nicostrato D., Kassim, Yumna, Ringler, Claudia, Thomas, Timothy S., ElDidi, Hagar, Breisinger, Clemens, 2021-09-24 The importance of a resilient agriculture sector in providing food security livelihoods and household income was highlighted in many countries by the recent pandemic as was the capacity of the sector to cushion the negative impacts of the subsequent economic slowdown This has been the case in Egypt where agriculture has been resilient to the health crisis in comparison with the service and industry sectors Breisinger et al 2020 However the sector's resiliency is gradually being corroded by climate change with lasting harmful effects for agriculture and food systems

**Study Guide to Accompany Basics for Chemistry** Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross

reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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**Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to

the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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