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Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

 Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to lodine. This results in smaller size of Fluorine than lodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

- Indicate whether the following properties increase or decrease from left to right across the periodic table.
 - a. atomic radius (excluding noble gases)

 Decreases
 - b. first ionization energy
 - c. electronegativity

Increases

Increases

What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius decreases down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

lonization energy increases from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (lonization Energy) increases.

Periodic Properties Lab Answer Key

R. V. Digumarthi

Periodic Properties Lab Answer Key:

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,2015 A compilation of permanent rules and executive orders Including updates filed with the Secretary of State as of August 15 2011
Cellular Automata And Complexity Stephen Wolfram,2018-03-08 Are mathematical equations the best way to model nature For many years it had been assumed that they were But in the early 1980s Stephen Wolfram made the radical proposal that one should instead build models that are based directly on simple computer programs Wolfram made a detailed study of a class of such models known as cellular automata and discovered a remarkable fact that even when the underlying rules are very simple the behaviour they produce can be highly complex and can mimic many features of what we see in nature And based on this result Wolfram began a program of research to develop what he called A Science of Complexity The results of Wolfram's work found many applications from the so called Wolfram Classification central to fields such as artificial life to new ideas about cryptography and fluid dynamics This book is a collection of Wolfram's original papers on cellular automata and complexity Some of these papers are widely known in the scientific community others have never been published before Together the papers provide a highly readable account of what has become a major new field of science with important implications for physics biology economics computer science and many other areas

The book delves into Periodic Properties Lab Answer Key. Periodic Properties Lab Answer Key is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and indepth insights into Periodic Properties Lab Answer Key, encompassing both the fundamentals and more intricate discussions.

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 - Chapter 4: Periodic Properties Lab Answer Key in Specific Contexts
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