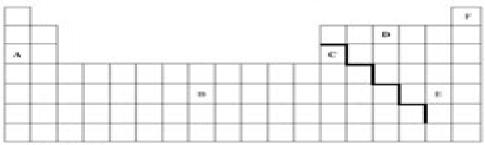
Part 1: Use the periodic table below to answer questions 1-12. Multiple answers may apply



- 1. Which letter(s) represent metals?
- 2. Which letter(s) represent nonmetals?
- Which of the following combinations represent an ionic compound?
 - a. A & B

c. A & E

c, C&F

b. B & E

d. C & D

- f. D & E
- Which of the following combinations represent a covalent molecule?
 - a. A & C

c. B & D

e. E & F

b. D & E

d. C&F

- f. C & D
- Which letter(s) represent elements that use roman numerals in their chemical name(s)?
- Which letter(s) will form a cation?
- 7. Which letter(s) will form an anion?
- Which letter(s) will gain three electrons to be stable?
- Which letter(s) will acquire a +3 charge?
- 10. Which letter has the highest ionization energy?
- 11. Which letter has the highest electronegativity?
- 12. Which letter(s) will acquire a positive charge in following the octet rule?

Part 2: Answer the questions below

- 13. How many electrons does Mg2+ have?
- 14. How many protons does N⁻³ have?
- 15. What will be the ion symbol (formula), if the ion has 19 protons and 18 electrons?
- 16. How many electrons does an ion have if it has 16 protons and a -2 charge? _____
- 17. If an ion has 53 protons and 54 electrons, what will be its charge?
- Given the formula X(NO₃)₃, what is the charge on ion X? Be sure to include the sign (+ or -).
- How many electrons does the Copper ion have in the ionic compound Cu₂SO₄?
- 20. Why do ionic compounds only in the liquid state conduct electricity?

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Barron's Educational Series, Mark Kernion, Joseph A. Mascetta

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