

Name: _____
Date: _____ Period: _____

Chemistry: *The Periodic Table*

Directions: Answer each of the following questions. You need not use complete sentences.

What is a period? How many are there in the periodic table?

A period is a row of elements on the periodic table, in which the elements are ordered one proton more each place from left to right. A period also often helps to represent an electron energy level and the order in which electrons are filled into their sublevels. There are 7 periods.

What is a group (also called a family)? How many are there in the periodic table?

Groups are the columns that are found top to bottom in the table. They show elements that have the same sublevel and number of electrons in their outer shell, valence electrons. Often elements in a group share any similar chemical properties because of this. There are 18 groups in the periodic table.

State the number of valence electrons in an atom of:

a. sulfur
6

b. calcium
2

c. chlorine
7

d. arsenic
5

Give the names and chemical symbols for the elements that correspond to these atomic numbers:

a. 10

Neon, Ne

b. 18

Argon, Ar

c. 36

Krypton, Kr

d. 90

Thorium, Th

List, by number, both the period and group of each of these elements.

	Symbol	Period	Group
a. beryllium	Be	2	2
b. iron	Fe	4	8
c. lead	Pb	6	14

Which of the following pairs of elements belong to the same period?

a. Na and Cl

b. Na and Li

c. Na and Cu

d. Na and Ne

Which of the following pairs of elements belong to the same group?

a. H and He

b. Li and Be

c. C and Pb

d. Ga and Ge

Would you expect strontium to be, chemically, more similar to calcium or rubidium and WHY?

Calcium, because they share a common amount of valence electrons. Valence electrons are the main contributor to chemical properties and how elements interact with other matter. Both strontium and calcium have two electrons on the the "s" sublevel, $5s^2$ and $4s^2$

What are the Group 1 elements called?

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web hess s law extra ws name answer key 1 find the ΔH for the reaction below given the following reactions and subsequent ΔH values
 $2 \text{ CO}_2 \text{ g} + \text{H}_2\text{O g} \rightarrow \text{C}_2\text{H}_2 \text{ g} + \frac{5}{2} \text{O}_2 \text{ g}$ $\Delta H = 2 \text{ h}_2 \text{ g} + \text{C}_2\text{H}_6 \text{ g}$ $\Delta H = 94.5 \text{ kJ}$
 $\text{H}_2\text{O g} + \text{H}_2 \text{ g} + \frac{1}{2} \text{O}_2 \text{ g}$ $\Delta H = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g} + \frac{7}{2} \text{O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2\text{O g}$ $\Delta H = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g}$

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web chemistry 120 hess s law worksheet calculate ΔH for the reaction $\text{C}_2\text{H}_4 \text{ g} + \text{H}_2 \text{ g} \rightarrow \text{C}_2\text{H}_6 \text{ g}$ from the following data
 $\text{C}_2\text{H}_4 \text{ g} + 3 \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 2 \text{ H}_2\text{O l}$ $\Delta H = 1411 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g} + \frac{7}{2} \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2\text{O l}$ $\Delta H = 1560 \text{ kJ}$
 $\text{H}_2 \text{ g} + \frac{1}{2} \text{ O}_2 \text{ g} \rightarrow \text{H}_2\text{O l}$ $\Delta H = 285 \text{ kJ}$
 mole

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web hess s law states that if a process can be expressed as the sum of two or more steps the enthalpy change for the overall process is the sum of the ΔH values for each step to use hess s law two principles must be understood one if an equation is reversed the sign of the ΔH value is also reversed two if an equation is multiplied by a

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web thus for reaction 9 2 in which 1 mole of H_2SO_4 and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}}$

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