

Periodic Trends – Atomic Mass

- The **atomic mass** describes the mass of each individual element.
- **Atomic mass** increases moving **down** each group and **right** across a period.

Instructions: In the periodic table below, draw two arrows representing the direction that atomic mass **increases** across groups and periods.

	1																	18
1		2											13	14	15	16	17	
2																		
3			3	4	5	6	7	8	9	10	11	12						
4																		
5																		
6																		
7																		

Instructions: In each question, identify the element with the **greatest** atomic mass:

- 1) Chlorine (Cl), Iodine (I), Bromine (Br), Fluorine (F) Iodine
- 2) Oxygen (O), Carbon (C), Nitrogen (N), Boron (B) Oxygen
- 3) Lithium (Li), Silicon (Si), Sulfur (Br), Xenon (Xe) Xenon

Instructions: In each question, list the elements from **least** to **greatest** atomic mass:

- 4) Magnesium (Mg), Beryllium (Be), Barium (Ba) Be, Mg, Ba
- 5) Gallium (Ga), Selenium (Se), Potassium (K) K, Ga, Se

Periodic Trends In Atomic Properties Answers

L Reisser



Periodic Trends In Atomic Properties Answers:

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