

# Periodic Trends Practice Worksheet

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Use the periodic table and your knowledge of the periodic trends to answer the following questions.

- ① Identify each element as a metal, metalloid, or nonmetal
  - Ⓐ germanium \_\_\_\_\_
  - Ⓑ zinc \_\_\_\_\_
  - Ⓒ phosphorous \_\_\_\_\_
  - Ⓓ lithium \_\_\_\_\_
- ② Which of the two species is larger in each of the following pairs?
  - Ⓐ  $\text{N}^{3-}$  or  $\text{F}^-$  \_\_\_\_\_
  - Ⓑ  $\text{Mg}^{2+}$  or  $\text{Ca}^{2+}$  \_\_\_\_\_
  - Ⓒ  $\text{Fe}^{2+}$  or  $\text{Fe}^{3+}$  \_\_\_\_\_
- ③ Which of the two species is smaller in each of the following pairs?
  - Ⓐ  $\text{K}^+$  or  $\text{Li}^+$  \_\_\_\_\_
  - Ⓑ  $\text{Au}^+$  or  $\text{Au}^{3+}$  \_\_\_\_\_
  - Ⓒ  $\text{P}^{3-}$  or  $\text{N}^{3-}$  \_\_\_\_\_
- ④ What atom has the LARGEST atomic radius?  
\_\_\_\_\_
- ⑤ What atom has the SMALLEST atomic radius?  
\_\_\_\_\_
- ⑥ Which is larger?
  - Ⓐ  $\text{Mg}$  or  $\text{Mg}^{2+}$  \_\_\_\_\_
  - Ⓑ  $\text{I}$  or  $\text{I}^-$  \_\_\_\_\_
  - Ⓒ  $\text{Cl}^-$  or  $\text{Br}^-$  \_\_\_\_\_
  - Ⓓ  $\text{F}^-$  or  $\text{Ne}$  \_\_\_\_\_
  - Ⓔ  $\text{N}^{3-}$  or  $\text{P}^{3-}$  \_\_\_\_\_
- ⑦ Put the following in order of increasing radius:  $\text{Br}^-$   $\text{Se}^{2-}$   $\text{Sr}^{2+}$   $\text{Kr}$   $\text{Rb}^+$
- ⑧ Circle the correct answer.
  - Ⓐ Bromine has a higher / lower ionization energy than Lead
  - Ⓑ Barium has a higher / lower ionization energy than Radon
  - Ⓒ Fluorine has a higher / lower ionization energy than Arsenic
  - Ⓓ Lithium has a higher / lower ionization energy than Cesium
- ⑨ Which of the two species is larger in each of the following pairs?
  - Ⓐ  $\text{Mg}$  or  $\text{Sr}$
  - Ⓑ  $\text{Be}$  or  $\text{N}$
  - Ⓒ  $\text{P}$  or  $\text{Cl}$
  - Ⓓ  $\text{Ni}$  or  $\text{Br}$
- ⑩ Arrange the following in order of increasing electronegativity:  $\text{S}$ ,  $\text{O}$ ,  $\text{Ne}$ ,  $\text{Al}$

# Periodic Trends Practice Answers

**Christopher R. Hren, John T.  
Moore, Peter J. Mikulecky**



## Periodic Trends Practice Answers:

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