

REACTION RATES & EQUILIBRIUM

- REACTION RATE = SPEED OF REACTION
 - TELLS YOU HOW MUCH OF THE REACTANTS ARE CONVERTED TO PRODUCTS OVER A PERIOD OF TIME.
- SOME REACTIONS ARE VERY SLOW, OTHERS ARE VERY FAST.
- COLLISION THEORY: IN ORDER TO COMBINE CHEMICALLY, THE REACTANT PARTICLES MUST MAKE CONTACT AND HAVE ENOUGH ENERGY.
 - W/O ENOUGH ENERGY → JUST BOUNCE APART
- ACTIVATION ENERGY = MINIMUM ENERGY COLLIDING PARTICLES MUST HAVE TO REACT. (BARRIER REACTANTS MUST PASS TO BECOME PRODUCTS)

Reaction Rates And Equilibrium Packet

Ernő Keszei



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