

Name: _____ per: _____

Worksheet - Reaction Rates

Use this reaction for the questions below: $C_6H_{12}O_6 (s) + 6 O_2(g) \rightarrow 6 H_2O (g) + 6 CO_2 (g)$

- What happens to the concentrations of:
 - $C_6H_{12}O_6$ & O_2 as the reaction proceeds \rightarrow ?
 - H_2O + CO_2 as the reaction proceeds \rightarrow ?
- According to the collision theory, what 3 circumstances are needed for $C_6H_{12}O_6$ & O_2 to react?
- What is the activation energy for a chemical reaction?
- Use the equation & the collision theory to explain:



Change in condition:	Does this <u>increase</u> or <u>decrease</u> the rate of reaction?	Explain why
a. <u>Increasing the temperature</u>	Ex: Increase (springs up)	Ex: Molecules move faster & collide more = increased rxn rate
b. <u>Increasing the concentration</u> of $C_6H_{12}O_6$		
c. <u>Decreasing the concentration</u> of O_2		
d. <u>Increase the surface area</u> by chewing up food in your mouth.		
e. <u>Decreasing the temperature</u>		
f. <u>Increasing the pressure</u> in the container		
g. <u>Decreasing the concentration</u> of H_2O		
h. <u>Increasing the volume</u> of the container the reaction occurs in		
i. <u>Increasing the concentration</u> of CO_2		
j. <u>Using a catalyst</u> (like salivary amylase)		

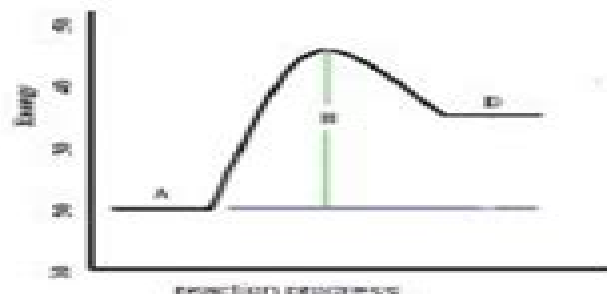
- On the accompanying energy diagram, label the following terms:

a. reactants b. products c. activation energy

- On the graph to the right, draw and label what this diagram would look like if a catalyst was added to the reaction.

- Graph reading

- How much energy (kJ) do the reactants have?
- How much energy (kJ) do the products have?
- How much energy (kJ) is required to activate this un-catalyzed reaction?
- Is this reaction endothermic or exothermic? How do you know?



Reaction Rates And Equilibrium Worksheet Answers

Regents Exambusters

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the 'Regents Exambusters' text.

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The Equilibrium Assumption in the Theory of Absolute Reaction Rates Charles Francis Curtiss, 1948 The assumption that systems in the activated state are in equilibrium with the reactant molecules has been open to question In this paper the methods of classical mechanics are used to obtain an expression for the rate constant without making this assumption The result is similar in form to that obtained by Eyring except that the average velocity of the points in configuration space is related directly to the distribution in velocities of the molecules of the gas in real space If one assumes that this distribution is Maxwellian one obtains the Eyring expression However the chemical reaction itself acts as a perturbing influence and causes the distribution in velocities to be somewhat different from Maxwellian It is this perturbation of the distribution function which leads to a different value of the rate constant and consequently to a measure of the effect of the equilibrium assumption Author 2003, *NY Regents Chemistry Test Prep Review--Exambusters Flashcards* Regents Exambusters, 2016-06-01 NY Regents CHEMISTRY Study Guide 700 questions and answers Essential definitions formulas concepts and sample problems Topics Introduction Matter Atoms Formulas Moles Reactions Elements Periodic Table Electrons Chemical Bonds Heat Gases Phase Changes Solutions Reaction Rates Equilibrium Acids and Bases Oxidation and Reduction Introduction to Organic Chemistry Radioactivity ADDITIONAL WORKBOOKS NY Regents INTEGRATED ALGEBRA Study Guide 450 questions and answers Essential definitions formulas concepts and sample problems Topics Sets Variables Exponents Properties of Numbers Like Terms Simple Equations Property of Equality Signed Numbers Monomials Polynomials Advanced Equations Verbal Problems Factoring Polynomials Algebraic Fractions Equations with Several Variables Advanced Verbal Problems Evaluating Formulas Simultaneous Equations Ratio and Proportion Variation Quadratic Equations and Radicals Coordinate Geometry _____ NY Regents UNITED STATES HISTORY Study Guide 700 questions and answers ILLUSTRATED Essential names dates and summaries of key historical events Topics Discovery Colonial Revolutionary Early National Age of Expansion

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