

**Unit 11: Chapter 7 - Chemical Reactions Review Answer Key**

**Define:**

1. **Reactant:** *A substance that undergoes change in a chemical reaction.*
2. **Product:** *New substances formed as a result of a chemical reaction.*
3. **Catalyst:** *A substance that affects the rate of a chemical reaction without being used up in the reaction.*
4. **Coefficient:** *Numbers that appear before a formula in a chemical equation to show the relative proportions of each reactant and product.*
5. **Subscript:** *A number written lower in a chemical formula indicating the amount of atoms present of an element (this number is not to be changed once written in chemical formula).*
6. **Law of Conservation of Mass:** *Matter cannot be created nor destroyed; only changed (atoms can be rearranged).*
7. **Synthesis reaction (include an example):** *A chemical reaction in which two or more substances react to form a single substance.*  $(A + B \rightarrow AB)$
8. **Decomposition reaction (include an example):** *A chemical reaction in which a compound breaks down into two or more simpler substances.*  $(AB \rightarrow A + B)$
9. **Single replacement (include an example):** *A chemical reaction in which one element takes the place of another element in a compound.*  $(AB + C \rightarrow AC + B)$
10. **Double replacement (include an example):** *A chemical reaction in which two compounds exchange positive ions and form two new compounds.*  $(AB + CD \rightarrow AD + CB)$
11. **Combustion (include an example):** *A chemical reaction in which a substance reacts rapidly with oxygen, often producing heat and light.*  $(C_2H_2 + O_2 \rightarrow CO_2 + H_2O)$
12. **Endothermic reaction:** *A chemical reaction that absorbs energy from its surroundings.*
13. **Exothermic reaction:** *A chemical reaction that releases energy to its surroundings.*
14. **Chemical equation (include an example):** *A representation of a chemical reaction in which the reactants and products are expressed as formulas.*
15. A **balanced equation** is an equation in which the number of atoms on the reactants' side **equals** the number of atoms on the products' side of the chemical equation.
16. **Equilibrium:** *A state in which the forward and reverse paths of a physical or chemical change take place at the same rate.*

# Reactions Review Answer

**Neil D. Jespersen, Alison Hyslop**



## Reactions Review Answer:

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Jespersen, Alison Hyslop, 2021-11-02 Chemistry The Molecular Nature of Matter 8th Edition continues to focus on the intimate relationship that exists between structure at the atomic molecular level and the observable macroscopic properties of matter Key revisions in this edition focus on three areas The deliberate inclusion of more updated real world examples that relate common real world student experiences to the science of chemistry Simultaneously examples and questions have been updated to align them with career concepts relevant to the environmental engineering biological pharmaceutical and medical sciences Providing students with transferable skills with a focus on integrating metacognition and three dimensional learning into the text When students know what they know they are better able to learn and incorporate the material Providing a total solution through New WileyPLUS by fully integrating the enhanced etext with online assessment answer specific responses and additional practice resources The 8th edition continues to emphasize the importance of applying concepts to problem solving to achieve high level learning and increase retention of chemistry knowledge Problems are arranged in an intuitive confidence building order

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