

3 • Reactions In Aqueous Solution**PRACTICE TEST**

- On the basis of the solubility rules, which of the following is insoluble?
a) K_2O d) $(NH_4)_2SO_4$
b) Na_2CO_3 c) $Ba(C_2H_3O_2)_2$
e) PbS
- In a double replacement reaction, formation of which of the following does not necessarily lead to a chemical change?
a) $HC_2H_3O_2$ d) H_2S
b) $AgCl$ e) $NaCl$
c) CO_2
- Reaction of an acid with a carbonate (such as $CaCO_3$) always results in the formation of
a) O_2 d) O_3
b) $C_{(diamond)}$ c) CO_2
e) CH_4
- Which of the following is incorrect?
a) all salts containing NH_4^+ are soluble.
b) all salts containing NO_3^- are soluble.
c) all fluorides are soluble.
d) all sulfates (except those of Ca^{2+} , Sr^{2+} , Ba^{2+} , and Pb^{2+}) are soluble.
e) most hydroxides are insoluble, except those of Ca^{2+} , Sr^{2+} , Ba^{2+} , the alkali metals and NH_4^+ .
- One of the gases shown below is NOT usually formed in a double replacement reaction. Which one?
a) N_2 d) NH_3
b) CO_2 c) H_2S
e) SO_2
- Write the balanced molecular equation for the reaction of washing soda, Na_2CO_3 and vinegar, $HC_2H_3O_2$.
- The net ionic equation for the above reaction is:
- How many moles of H^+ are associated with the acid, H_2SO_3 , during neutralization?
a) 0 b) 1 c) 2 d) 3
- How many moles Al_2O_3 are needed to neutralize 1 mole of HCl ?
a) $1/3$ d) 6
b) $2/3$ e) 12
c) 2 f) $1/6$
- Write the net reaction that will occur when solid ammonium carbonate is added to a solution of hydrosulfuric acid.

Reactions In Aqueous Solution Practice Test

Gary S Thorpe



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