

# Reactivity Series

Metal (Ion)	Reactivity	Most reactive
Potassium (K <sup>+</sup> )	Reacts with water	
Sodium (Na <sup>+</sup> )		
Lithium (Li <sup>+</sup> )		
Barium (Ba <sup>2+</sup> )		
Strontium (Sr <sup>2+</sup> )		
Calcium (Ca <sup>2+</sup> )		
Magnesium (Mg <sup>2+</sup> )	Reacts with acid	
Aluminum (Al <sup>3+</sup> )		
Manganese (Mn <sup>+</sup> )		
Zinc (Zn <sup>2+</sup> )		
Chromium (Cr <sup>2+</sup> )		
Iron (Fe <sup>2+</sup> )		
Cadmium (Cd <sup>2+</sup> )		
Cobalt (Co <sup>2+</sup> )		
Nickle (Ni <sup>2+</sup> )		
Tin (Sn <sup>2+</sup> )		
Lead (Pb <sup>2+</sup> )		
Hydrogen (H <sup>+</sup> )	Included for comparison	
Antimony (Sb <sup>2+</sup> )	Highly unreactive	
Bismuth (Bi <sup>2+</sup> )		
Copper (Cu <sup>2+</sup> )		
Mercury (Hg <sup>2+</sup> )		
Silver (Ag <sup>+</sup> )		
Gold (Au <sup>3+</sup> )		
Platinum (Pt <sup>+</sup> )	Least reactive	

# Reactivity Series Pogil

**Carol Baldwin**



## Reactivity Series Pogil:

Reactivity Series & Displacement Reactions, 2006      **Reactivity Series**, 2006      *Chemical Reactions, Materials and Particles* Gill Murphy, 2006-03 Specials are a teaching resource designed for KS3 students whose literacy skills are considerably lower than their age. These books have an older format to counteract the simple text and cover various topics. They include activities, visuals, and assessment sheets as well as teacher pages.      Mechanism and Reactivity for a Series of Hydrolytic Reactions Terry Carl Pletcher, 1969      *Material Changes and Reactions* Chris Oxlade, 2007. What happens in a redox reaction? How do materials change when they melt or freeze? Why are chemical reactions so important? This title explores how elements melt, boil, and freeze; how they combine with each other to make new substances; and how chemical reactions take place every day in factories and in your home. You will also find several experiments that can be done at home.

**Chemical Reactions** Simon de Pinna, 2007-01-12. Introduces chemical reactions covering how atoms, compounds, and mixtures form acids, bases, and salts, and what happens when these come in contact with each other under certain conditions.

Chemical Reactions Carol Baldwin, 2006. Describes the different types of chemical reactions and how temperature, concentration, particle size, and catalysts affect the reaction rate.      **Solvent Effects and Chemical Reactivity** Orlando Tapia, Juan Bertrán, 2006-04-11. This book gathers original contributions from a selected group of distinguished researchers that are actively working in the theory and practical applications of solvent effects and chemical reactions. The importance of getting a good understanding of surrounding media effects on chemical reacting systems is difficult to overestimate. Applications go from condensed phase chemistry, biochemical reactions in vitro to biological systems in vivo. Catalysis is a phenomenon produced by a particular system interacting with the reacting subsystem. The result may be an increment of the chemical rate or sometimes a decreased one. At the bottom, catalytic sources can be characterized as a special kind of surrounding medium effect. The materials involving in catalysis may range from inorganic components as in zeolites, homogenous components, enzymes, catalytic antibodies, and ceramic materials. With the enormous progress achieved by computing technology, an increasing number of models and phenomenological approaches are being used to describe the effects of a given surrounding medium on the electronic properties of selected subsystems. A number of quantum chemical methods and programs currently applied to calculate in vacuum systems have been supplemented with a variety of model representations. With the increasing number of methodologies applied to this important field, it is becoming more and more difficult for non-specialists to cope with theoretical developments and extended applications. For this and other reasons, it was deemed timely to produce a book where methodology and applications were analyzed and reviewed by leading experts in the field.      **The Reaction Path in Chemistry: Current Approaches and Perspectives** D. Heidrich, 2013-03-09. The so-called reaction path (RP) with respect to the potential energy or the Gibbs energy/free enthalpy is one of the most fundamental concepts in chemistry. It significantly helps to display and visualize the results of the complex microscopic processes forming

a chemical reaction This concept is an implicit component of conventional transition state theory TST The model of the reaction path and the TST form a qualitative framework which provides chemists with a better understanding of chemical reactions and stirs their imagination However an exact calculation of the RP and its neighbourhood becomes important when the RP is used as a tool for a detailed exploring of reaction mechanisms and particularly when it is used as a basis for reaction rate theories above and beyond TST The RP is a theoretical instrument that now forms the theoretical heart of direct dynamics It is particularly useful for the interpretation of reactions in common chemical systems A suitable definition of the RP of potential energy surfaces is necessary to ensure that the reaction theories based on it will possess sufficiently high quality Thus we have to consider three important fields of research Analysis of potential energy surfaces and the definition and best calculation of the RPs or at least of a number of selected and chemically interesting points on it The further development of concrete versions of reaction theory beyond TST which are applicable for common chemical systems using the RP concept

Chemical Reactions Louise Spilbury, Richard Spilbury, 2007 Readers will learn what chemical reactions are how they work what changes happen during reactions and how we can stop reactions

**Reactivity Patterns in a Series of Novel Monothiocarbamates, Dithiocarbamates, and Sulfur-rich Dithiolates** Donald M. Baird, 1982

*Chemistry and Chemical Reactivity* John C. Kotz, Keith F. Purcell, 1987

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