

1. Hydrogen gas and Iron chloride are produced.

Explanation: The Chlorine from Hydrogen chloride is displaced by the Iron fillings to undergo the following reaction.

2HCl + Fe → FeCl<sub>2</sub>+ H<sub>2</sub>

4. What is a balanced chemical equation? Why should a chemical equation be balanced?

#### Sobation:

A balanced equation is the one in which number of different atoms on both the reactant and product sides are equal. Balancing chemical equations is necessary for the reaction should obey The Law of Conservation of mass. Balancing the chemical equation has no defined method and is purely a trial-and-error attempt.

- 5. Translate the following statements into chemical equations and balance them.
- (a) Hydrogen gas combines with nitrogen to form ammonia.
- (b) Hydrogen sulphide gas burns in the air to give water and sulphur dioxide.

(c) Barium chloride reacts with aluminium sulphate to give Aluminium chloride and a precipitate of barium sulphate.

(d) Potassium metal reacts with water to give potassium hydroxide and Hydrogen gas.

#### Sobuttions:

- (a) Unbalanced: H<sub>2</sub> + N<sub>2</sub> → NH<sub>3</sub>
- Balanced: 3H.+ N. -- 2NH.
- (b) Unbalanced: H<sub>2</sub>S + O<sub>2</sub> → H<sub>2</sub>O + SO<sub>3</sub>
- Balanced:  $2H.S + 3O. \rightarrow 2H.O + 2SO.$
- (c) Unbalanced:
- $BaCl_t + Al_t(SO_t) \rightarrow AlCl_t + BaSO_t$
- Balanced:  $3BaCl_1 + Al_2(SO_4) \rightarrow 2AlCl_4 + 3BaSO_4$
- (d) Unbalanced: K + H<sub>2</sub>O → KOH + H<sub>3</sub>
- Balanced: 2K + 2H<sub>2</sub>O → 2KOH + H<sub>2</sub>

#### Balance the following chemical equations.

- (a)  $HNO_2 + Ca(OH)_1 \rightarrow Ca(NO_2)_2 + H_2O$
- (b)  $NnOH + H_iSO_i \rightarrow Nn_iSO_i + H_iO$
- (c) NaCl + AgNO<sub>2</sub> → AgCl + NaNO<sub>2</sub>
- (d) BaCl<sub>2</sub> + H<sub>2</sub>SO<sub>4</sub> → BaSO<sub>4</sub> + HCl

#### Solution:

(a) 2HNO<sub>i</sub>+ Ca(OH)<sub>i</sub>→ Ca(NO<sub>i</sub>)<sub>i</sub>+ 2H<sub>i</sub>O

# **Section 1 Chemical Changes Answers**

**Amar Bhutani, Priya Minhas** 

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## **Table of Contents Section 1 Chemical Changes Answers**

- 1. Understanding the eBook Section 1 Chemical Changes Answers
  - The Rise of Digital Reading Section 1 Chemical Changes Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Section 1 Chemical Changes Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Section 1 Chemical Changes Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Section 1 Chemical Changes Answers
  - Personalized Recommendations
  - Section 1 Chemical Changes Answers User Reviews and Ratings
  - Section 1 Chemical Changes Answers and Bestseller Lists
- 5. Accessing Section 1 Chemical Changes Answers Free and Paid eBooks
  - Section 1 Chemical Changes Answers Public Domain eBooks
  - Section 1 Chemical Changes Answers eBook Subscription Services
  - Section 1 Chemical Changes Answers Budget-Friendly Options
- 6. Navigating Section 1 Chemical Changes Answers eBook Formats

- o ePub, PDF, MOBI, and More
- Section 1 Chemical Changes Answers Compatibility with Devices
- Section 1 Chemical Changes Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - o Adjustable Fonts and Text Sizes of Section 1 Chemical Changes Answers
  - Highlighting and Note-Taking Section 1 Chemical Changes Answers
  - Interactive Elements Section 1 Chemical Changes Answers
- 8. Staying Engaged with Section 1 Chemical Changes Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Section 1 Chemical Changes Answers
- 9. Balancing eBooks and Physical Books Section 1 Chemical Changes Answers
  - Benefits of a Digital Library
  - o Creating a Diverse Reading Collection Section 1 Chemical Changes Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Section 1 Chemical Changes Answers
  - Setting Reading Goals Section 1 Chemical Changes Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 1 Chemical Changes Answers
  - Fact-Checking eBook Content of Section 1 Chemical Changes Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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