

7-1 Review and Reinforcement

Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative
positive
crisscross

noble gas
electrons
anion

octet
ions
cation

electron configuration
polyatomic
empirical formula

1. An ionic bond is an attraction between oppositely charged ions.
2. Anions have a negative charge.
3. An atom becomes an ion by losing or gaining electrons.
4. The octet rule states that atoms tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of a(n) noble gas.
6. The atoms that make up polyatomic ions are bonded together by covalent bonds.
7. The crisscross method can be used to write the formula for an ionic compound.
8. The empirical formula of a compound denotes the ratio of ions in the compound.

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line.

- | | |
|------------------------|--------------------------------------------------------------------------------------------|
| <u>losing</u> | 9. Calcium becomes a monatomic cation by <u>gaining</u> two electrons. |
| <u>true</u> | 10. A cation has a <u>positive</u> charge. |
| <u>true</u> | 11. A binary ionic compound contains only <u>one</u> kind of cation and one kind of anion. |
| <u>polyatomic</u> | 12. <u>Monatomic</u> ions consist of more than one atom. |
| <u>seven</u> | 13. The Lewis dot diagram for chlorine, a group 7A element, has <u>six</u> electrons. |
| <u>true (released)</u> | 14. A great deal of energy is <u>produced</u> when an ionic compound is formed. |

Section 1 Reinforcement Atoms Answers

J Elliott



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Table of Contents Section 1 Reinforcement Atoms Answers

1. Understanding the eBook Section 1 Reinforcement Atoms Answers
 - The Rise of Digital Reading Section 1 Reinforcement Atoms Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 1 Reinforcement Atoms Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 1 Reinforcement Atoms Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 1 Reinforcement Atoms Answers
 - Personalized Recommendations
 - Section 1 Reinforcement Atoms Answers User Reviews and Ratings
 - Section 1 Reinforcement Atoms Answers and Bestseller Lists
5. Accessing Section 1 Reinforcement Atoms Answers Free and Paid eBooks

- Section 1 Reinforcement Atoms Answers Public Domain eBooks
- Section 1 Reinforcement Atoms Answers eBook Subscription Services
- Section 1 Reinforcement Atoms Answers Budget-Friendly Options
- 6. Navigating Section 1 Reinforcement Atoms Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Section 1 Reinforcement Atoms Answers Compatibility with Devices
 - Section 1 Reinforcement Atoms Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 1 Reinforcement Atoms Answers
 - Highlighting and Note-Taking Section 1 Reinforcement Atoms Answers
 - Interactive Elements Section 1 Reinforcement Atoms Answers
- 8. Staying Engaged with Section 1 Reinforcement Atoms Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 1 Reinforcement Atoms Answers
- 9. Balancing eBooks and Physical Books Section 1 Reinforcement Atoms Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 1 Reinforcement Atoms Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section 1 Reinforcement Atoms Answers
 - Setting Reading Goals Section 1 Reinforcement Atoms Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 1 Reinforcement Atoms Answers
 - Fact-Checking eBook Content of Section 1 Reinforcement Atoms Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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