

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from pH and pOH , and vice versa from pH and pOH .
3. Calculate the pH of a weak acid.

Vocabulary

- | | | |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. pH | 6. acid solution |

Key Equation

- | | |
|------------------------------------|------------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = -\log[\text{H}^+]$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = -\log[\text{H}^+]$ |

Part A Completion

Write the hydrogen ion concentration or autoionization constant, if the missing quantity that is indicated in the partial equation below. Complete each equation as needed.

1. What substance has _____ autoionization constant?
2. _____ is the autoionization of water which produces an H^+ ion and a hydroxide ion.
3. _____ is the autoionization of water which produces _____ and _____.
4. _____ is the autoionization of water which produces _____ and _____.
5. _____ is the autoionization of water which produces _____ and _____.
6. _____ is the autoionization of water which produces _____ and _____.
7. _____ is the autoionization of water which produces _____ and _____.
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18. _____ is the autoionization of water which produces _____ and _____.
19. _____ is the autoionization of water which produces _____ and _____.
20. _____ is the autoionization of water which produces _____ and _____.

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Section 19 2 Hydrogen And Ions Acidity Answer Key

TD Snyder



Section 19 2 Hydrogen And Ions Acidity Answer Key:

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