OMPTER (2) STUDY GUIDE FOR CONTENT MASTERS

Section 12.3 Limiting Reactants

in your textbook, read about any mactions stop and how to determine the Uniting reaction.

Study the diagram showing a chemical reaction and the elemical equation that represents the traction. Then complete the table. Show your calculations for questions 25-27 in the space below the table.

$$\Box + 00 + 0 \rightarrow 0 + 00$$

 $0,+200 \rightarrow 200$

The media masses of O₂, NO, and NO₂ are 32.00 g/mel, 30.01 g/mel, and 46.01 g/mel, emporatedly.

Amount of O ₂	Annual of NO	Amount of Mily	Similing Number	Amount and Name of Quees Resident
1 molecule	Teolesies	2 molecules	mine	none
Challetales	4 mileciles	4 molecules	NO	2 milesia 0 _y
I molecules	1 militaire	1. 4 molecules	2. 02	1, 4 molecules NO
3.88 mol -	2.00 mail	4. 2.00 mol	5. none	6. none
438 mol	450 mail	7, 4.00 mail	E. NO	9, 2.00 mol 0 ₂
3.00 mail	7.00 mult	16, 7.00 mol	11, NO	12, 1.50 mail 0 ₂
1.00 mol	400 mil	13. 2.00 mol	14. O ₃	15. 2.00 mol NO
1.500 mail	0.200 mail	16, 0,200 mol	17, 100	18. 0.400 mol O ₂
name	mag	19. 92.02 g	20. none	21. none
26.00 (6	mong	22, 46.01 g	20. 02	34, 50.12 g NO
mm ₁	20.00 g	25, 28.74 g	26. 02	37, 1,34 g NO

balanced equation mole ratio = 2 mol NO/1 mol O2

10.00 g.O₂ × 1 mol O₂/32.00 g.O₂ = 0.3125 mol O₂

20.00 p.NO × 1 mai NO/30.01 p.NO = 0.6664 mai NO

actual mole ratio = 0.6664 mol NO/0.3125 mol O₂ = 2.132 mol NO/1.000 mol O₂ Secause the actual mole ratio of NO:O₂ is larger than the balanced equation mole ratio of NO:O₃, there is an excess of NO; O₃ is the limiting reactant.

Mass of NO used = 0.3125 gpai-0; × 2 mol NO/1 gpai-0; = 0.6250 mol NO 0.6250 gpai-NO × 30.01 g NO/1 gpai-NO = 18.76 g NO

Mass of NO₂ produced = 0.6250 gpsi.WO₂ × 46.01 g NO₂/1 gpsi.WO₂ = 28.76 g NO₂ Excess NO = 20.00 g NO - 18.76 g NO = 1.24 g NO

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Chemistry Matter and Change + Chapter 12 ... 29.



Section 12.4 Percent Yield

In your textlisiol, resel about the pielos of products.

Study the diagram and the example problem.

mass of product from experimental measurement

mais of product predicted from assistionietric calculation using

- at mass of reactions.
- 5. 4-dep main to main conversion
 - 1. Sems the belanced chemical equation.
 - Calculate the number of moles of reactions, using moles mass.
 - Calculate the number of moles of product, using the appropriate mole ratio.
 - Calculate the mass of product, using the reciprocal of molar mass.

Example Problem: The following chemical equation represents the production of gallium oriole, a substance used in the manufacturing of some semiconductor devices.

In one experiment, the reaction yielded 7.42 g of the enals from a 7.85 g sample of gallium. Describes the percent yield of this reaction. The molar masses of Ga and Ga₂O₂ are 46.72 glossl and 167.46 glossl, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- 1. actual right unknown 7.42 g Ga₂O₃
- 2. mass of touchast: Total y the connect.
- matter of motion of reactant: 7.00 g/G₀ × min(T₂ g/G₀).
 mod G₀.
 mod G₀.
 mod G₀.
 mod G₀.
- 4. number of moles of product 0.000 mol $G_0 \times \frac{2 \bmod G_0 / G_0}{1 \bmod G_0} = \frac{0.100 \bmod G_0 \times 2 \bmod G_0}{2 \bmod G_0 / 2 \bmod G_0}$
- 5. theoretical yield. 0:0500 and $O_{m_0}O_{\gamma} \times \frac{107.44 \pm O_{m_0}O_{\gamma}}{1 \text{ and } O_{m_0}O_{\gamma}}$ connect.
- percent yield: 41.7 g Ga₂O₃ × 100
 7.42 g Ga₂O₃/9.37 g Ga₂O₃ × 100
- 22 Chemistry, Matter and Change + Chapter 12

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