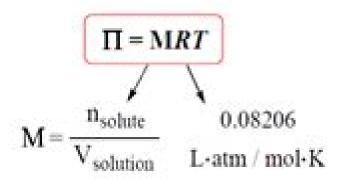
## **Key Equations for Colligative Properties**

Raoult's Law: Relationship between the Vapor Pressure of a Solution ( $P_{\text{solution}}$ ), the Mole Fraction of the Solvent ( $\chi_{\text{solvent}}$ ), and the Vapor Pressure of the Pure Solvent ( $P_{\text{solvent}}$ )

The Vapor Pressure of a Solution Containing Two Volatile Components

$$P_{A} = \chi_{A} P^{o}_{A}$$
  $P_{B} = \chi_{B} P^{o}_{B}$  
$$P_{total} = P_{A} + P_{B}$$

Relationship between Osmotic Pressure  $(\Pi)$ , Molarity (M), the Ideal Gas Constant (R), and Temperature (T, in K).



The dependance of  $\Delta T_f$  and  $\Delta T_b$  on the Molality (m), and Freezing Point Depression Constant ( $K_f$ ) and Boiling Point Elevation Constant ( $K_b$ ) respectively.

$$\Delta T_f = m \cdot K_f$$

$$\Delta T_b = m \cdot K_b$$

$$m = \frac{\text{mol (solute)}}{\text{kg (solvent)}}$$

# **Section 18 3 Colligative Properties Of Solutions Answers**

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## **Section 18 3 Colligative Properties Of Solutions Answers:**

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