

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on  $K_a$  and  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                        |                             |                   |
|------------------------|-----------------------------|-------------------|
| 1. autoionization      | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization      | 5. $K_a$                    | 6. $K_b$          |
| 7. ionization constant | 8. $K_a$                    | 9. $K_b$          |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the equation and/or the numerical value(s) of the constants indicated that are associated in pH-related calculations with hydrogen ions and hydroxide ions.

1. What substance has  $\text{pH} = 7$ , contains hydrogen ions?
2. What substance has  $\text{pH} = 7$ , the concentration of which is the product of  $[\text{H}^+]$  and  $[\text{OH}^-]$ ?
3. What substance contains the highest concentration of  $\text{H}^+$ ?
4. What substance contains the highest concentration of  $\text{OH}^-$ ?
5. What substance contains the highest concentration of  $\text{H}^+$ ?
6. What substance contains the highest concentration of  $\text{OH}^-$ ?
7. What substance contains the highest concentration of  $\text{H}^+$ ?
8. What substance contains the highest concentration of  $\text{OH}^-$ ?
9. What substance contains the highest concentration of  $\text{H}^+$ ?
10. What substance contains the highest concentration of  $\text{OH}^-$ ?

1.  $\text{H}_2\text{O}$
2.  $1.0 \times 10^{-7} \text{ M}$
3.  $1.0 \times 10^{-7} \text{ M}$
4.  $1.0 \times 10^{-7} \text{ M}$
5.  $1.0 \times 10^{-7} \text{ M}$
6.  $1.0 \times 10^{-7} \text{ M}$
7.  $1.0 \times 10^{-7} \text{ M}$
8.  $1.0 \times 10^{-7} \text{ M}$
9.  $1.0 \times 10^{-7} \text{ M}$
10.  $1.0 \times 10^{-7} \text{ M}$

## Section 19 2 Hydrogen Ions And Acidity Answer Key

**CO Houle**



## **Section 19 2 Hydrogen Ions And Acidity Answer Key:**

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