

7-1 Review and Reinforcement

Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative
positive
electron

noble gas
atoms
ion

cation
anion
cation

electron configuration
polyatomic
empirical formula

1. An ionic bond is an attraction between oppositely charged _____. **ions**
2. Anionic bond is _____. **negative** change.
3. An ionic bond is formed by losing or gaining _____. **electrons**
4. The _____ noble gases that cannot bond to gain, lose, or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form a bond, each has negative the electron configuration of _____ noble gas.
6. The atoms that make up _____ are ion bonded together by ionic bonds.
7. The _____ method can be used to write the formula for an ionic compound.
8. The _____ empirical formula is a compound denotes the ratio of ions in the compound. If the element is iron, write "Fe" if it is below. Change the subscript until an equals to make it true. Write your answer on the line.

_____ **losing**

_____ **iron**

_____ **iron**

_____ **polyatomic**

_____ **cations**

_____ **iron (II) oxide**

9. Calcium boride is chemically written as _____ **gaining** two electrons.

10. A cation has a _____ **positive** charge.

11. A binary ionic compound contains only _____ **one** kind of cation and one kind of anion.

12. Substances that consist of more than one atom.

13. The Lewis dot diagram for chlorine, a group 7A element, has six _____ **electrons**.

14. A great deal of energy is _____ **produced** when an ionic compound is formed.

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S Ben Porath



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Table of Contents Section 1 Reinforcement Landforms Worksheet Answers

1. Understanding the eBook Section 1 Reinforcement Landforms Worksheet Answers
 - The Rise of Digital Reading Section 1 Reinforcement Landforms Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 1 Reinforcement Landforms Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 1 Reinforcement Landforms Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 1 Reinforcement Landforms Worksheet Answers
 - Personalized Recommendations
 - Section 1 Reinforcement Landforms Worksheet Answers User Reviews and Ratings

- Section 1 Reinforcement Landforms Worksheet Answers and Bestseller Lists
- 5. Accessing Section 1 Reinforcement Landforms Worksheet Answers Free and Paid eBooks
 - Section 1 Reinforcement Landforms Worksheet Answers Public Domain eBooks
 - Section 1 Reinforcement Landforms Worksheet Answers eBook Subscription Services
 - Section 1 Reinforcement Landforms Worksheet Answers Budget-Friendly Options
- 6. Navigating Section 1 Reinforcement Landforms Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Section 1 Reinforcement Landforms Worksheet Answers Compatibility with Devices
 - Section 1 Reinforcement Landforms Worksheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 1 Reinforcement Landforms Worksheet Answers
 - Highlighting and Note-Taking Section 1 Reinforcement Landforms Worksheet Answers
 - Interactive Elements Section 1 Reinforcement Landforms Worksheet Answers
- 8. Staying Engaged with Section 1 Reinforcement Landforms Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 1 Reinforcement Landforms Worksheet Answers
- 9. Balancing eBooks and Physical Books Section 1 Reinforcement Landforms Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 1 Reinforcement Landforms Worksheet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section 1 Reinforcement Landforms Worksheet Answers
 - Setting Reading Goals Section 1 Reinforcement Landforms Worksheet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 1 Reinforcement Landforms Worksheet Answers
 - Fact-Checking eBook Content of Section 1 Reinforcement Landforms Worksheet Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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