

DESCRIBING CHEMICAL REACTIONS

Section Review - Objectives

- Explain how to write a word equation
- Describe how to write a skeleton equation
- List the steps for writing a complete chemical equation

Vocabulary

- chemical equation
- skeleton equation
- catalyst
- coefficients
- balanced equation

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

A chemical reaction can be concisely represented by a chemical ____1____. The substances that undergo a chemical change are the ____2____. The new substances formed in a chemical reaction are the ____3____. In accordance with the law of conservation of ____4____, a chemical equation must be balanced. When balancing an equation, you place ____5____ in front of reactants and products so that the same number of atoms of each ____6____ are on each side of the equation. An equation must never be balanced by changing the ____7____ in the chemical formula of a substance.

Special symbols are used to show the physical state of a substance in a reaction. The symbol for a liquid is ____8____; for a solid, ____9____; and for a gas, ____10____. A substance dissolved in water is designated ____11____. If a ____12____ is used to increase the rate of a chemical reaction, its formula is written above the arrow.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____
11. _____
12. _____

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- _____ 13. In an equation, a substance is shown to be in the gaseous state by placing an upward-pointing arrow after its formula.
- _____ 14. The symbol Δ placed over the arrow in an equation means that heat is supplied to the reaction.
- _____ 15. Atoms are destroyed in a chemical reaction.
- _____ 16. A skeleton equation is not a balanced equation.

Part D Questions and Problems

24. Write a balanced equation for each of these chemical reactions. I
- a. Aluminum reacts with aqueous hydrochloric acid to form hydrogen gas and aqueous aluminum chloride.
 - b. Acetylene gas (C_2H_2) burns in a welding torch with oxygen to form carbon dioxide gas and water vapor.

Section 1describing Chemical Reactions Answers

Kais A. M. Al Naimee



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Essentials of Chemical Reaction Engineering H. Scott Fogler, 2011 Accompanying DVD ROM contains many realistic interactive simulations The Practice of Chemistry Donald J. Wink, Sharon Fetzner-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink http://www.bfwpub.com/pdfs/wink/POC/PowerPoint_Final.ppt 832KB **Cehmistry Textbook for College and University USA** Ibrahim Sikder, 2023-06-04 Cehmistry Textbook USA **Principles of Modern Chemistry** David W. Oxtoby, H. Pat Gillis, Laurie J. Butler, 2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom **Qualitative Theory of Parabolic Equations, Part 1** T. I. Zelenyak, M. P. Vishnevskii, M. M. Lavrentiev, 2011-09-06 No detailed description available for Qualitative Theory of Parabolic Equations Part 1 Numerical Solution of Time-Dependent Advection-Diffusion-Reaction Equations Willem Hundsdorfer, Jan G. Verwer, 2013-04-17 This book deals with numerical methods for solving partial differential equations PDEs coupling advection diffusion and reaction terms with a focus on time dependency A combined treatment is presented of methods for hyperbolic problems thereby emphasizing the one way wave equation methods for parabolic problems and methods for stiff and non stiff ordinary differential equations ODEs With regard to time dependency we have attempted to present the algorithms and the discussion of their properties for the three different types of differential equations in a unified way by using semi discretizations i.e. the method of lines whereby the PDE is transformed into an ODE by a suitable spatial discretization In addition for hyperbolic problems we also discuss discretizations that use information based on characteristics Due to this combination of methods this book differs substantially from more specialized textbooks that deal exclusively with numerical methods for either PDEs or ODEs We treat integration methods suitable for both classes of

problems This combined treatment offers a clear advantage On the one hand in the field of numerical ODEs highly valuable methods and results exist which are of practical use for solving time dependent PDEs something which is often not fully exploited by numerical PDE researchers Although many problems can be solved by Euler s method or the Crank Nicolson method better alternatives are often available which can significantly reduce the computational effort needed to solve practical problems

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Warnant,2025-10-20 **Open-Ended Problems** James Patrick Abulencia,Louis Theodore,2015-03-27 This is a unique book with nearly 1000 problems and 50 case studies on open ended problems in every key topic in chemical engineering that helps to better prepare chemical engineers for the future The term open ended problem basically describes an approach to the solution of a problem and or situation for which there is not a unique solution The Introduction to the general subject of open ended problems is followed by 22 chapters each of which addresses a traditional chemical engineering or chemical engineering related topic Each of these chapters contain a brief overview of the subject matter of concern e g thermodynamics which is followed by sample open ended problems that have been solved by the authors employing one of the many possible approaches to the solutions This is then followed by approximately 40 45 open ended problems with no solutions although many of the authors solutions are available for those who adopt the book for classroom or training purposes A reference section is included with the chapter s contents Term projects comprised of 12 additional chapter topics complement the presentation This book provides academic industrial and research personnel with the material that covers the principles and applications of open ended chemical engineering problems in a thorough and clear manner Upon completion of the text the reader should have acquired not only a working knowledge of the principles of chemical engineering but also and more importantly experience in solving open ended problems What many educators have learned is that the applications and implications of open ended problems are not only changing professions but also are moving so fast that many have not yet grasped their tremendous impact The book drives home that the open ended approach will revolutionize the way chemical engineers will need to operate in the future Atomic Layer Deposition of Nanostructured Materials Nicola Pinna,Mato Knez,2012-09-19 Atomic layer deposition formerly called atomic layer epitaxy was developed in the 1970s to meet the needs of producing high quality large area fl at displays with perfect structure and process controllability Nowadays creating nanomaterials and producing nanostructures with structural perfection is an important goal for many applications in nanotechnology As ALD is one of the important techniques which offers good control over the surface structures created it is more and more in the focus of scientists The book is structured in such a way to fi t both the need of the expert reader due to the systematic presentation of the results at the forefront of the technique and their applications and the ones of students and newcomers to the fi eld through the first part detailing the basic aspects of the technique This book is a must have for all Materials Scientists Surface Chemists Physicists and Scientists in the Semiconductor Industry ODE, BVP, and 1D PDE Solvers for Scientific and Engineering Problems With MATLAB Basics Burstein, Leonid,2025-02-06 In the academic field engineers scientists educators and students are faced with a persistent challenge the gap between theoretical knowledge and practical implementation in solving real world engineering problems The scarcity of focused resources tailored to mastering MATLAB and its specialized solvers for Ordinary Differential Equations ODEs and One Dimensional Partial Differential Equations 1D PDEs has left many individuals struggling to bridge

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Chemical Misconceptions Keith Taber,2002 Part 2 provides strategies for dealing with some of the misconceptions that students have by including ready to use classroom resources

NASA Tech Briefs ,1995

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