

20

ACIDS AND BASES
PRACTICE PROBLEMS

In your notebook, solve the following problems.

SECTION 20.1 DESCRIBING ACIDS AND BASES

- Name the following compounds as acids.
 - HNO_2
 - H_2SO_4
 - HF
 - H_2CO_3
- Write the formulas for the following bases.
 - calcium hydroxide
 - ammonium hydroxide
 - aluminum hydroxide
 - lithium hydroxide
- Describe some distinctive properties of acids.
- Describe some distinctive properties of bases.

SECTION 20.2 HYDROGEN IONS AND ACIDITY

- A solution has a hydrogen ion concentration of $1 \times 10^{-6}\text{M}$. What is its pH?
- What is the pH of a solution if the $[\text{H}^+] = 7.2 \times 10^{-9}\text{M}$?
- What is the pOH of a solution if the $[\text{OH}^-] = 3.5 \times 10^{-3}\text{M}$?
- What is the pOH of a solution that has a pH of 3.4?
- Classify each solution as acidic, basic, or neutral.
 - $[\text{H}^+] = 2.5 \times 10^{-9}\text{M}$
 - pOH = 12.0
 - $[\text{OH}^-] = 9.8 \times 10^{-11}\text{M}$
 - $[\text{H}^+] = 1 \times 10^{-2}\text{M}$
 - pH = 0.8
- Calculate the pH of each solution.
 - $[\text{H}^+] = 1 \times 10^{-5}\text{M}$
 - $[\text{H}^+] = 4.4 \times 10^{-12}\text{M}$
 - $[\text{OH}^-] = 2.2 \times 10^{-7}\text{M}$
 - pOH = 1.4
- Classify the solutions in problem 6 as acidic or basic.
- Why is there a minus sign in the definition of pH?
- A solution has a pOH of 12.4. What is the pH of this solution?
- What is the pH of a solution with $[\text{H}^+] = 1 \times 10^{-2}\text{M}$?

SECTION 20.3 ACID-BASE THEORIES

- Identify the hydrogen ion donor(s) and hydrogen ion acceptor(s) for ionization of H_2SO_4 in water. Label the conjugate acid-base pairs.
- Identify all of the ions that may be formed when H_3PO_4 ionizes in water.

Section 20 Describing Acids And Bases Answers

Daniel C. Harris



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AP Chemistry For Dummies Peter J. Mikulecky, Michelle Rose Gilman, Kate Brutlag, 2008-11-13 A practical and hands on guide for learning the practical science of AP chemistry and preparing for the AP chem exam Gearing up for the AP Chemistry exam AP Chemistry For Dummies is packed with all the resources and help you need to do your very best Focused on the chemistry concepts and problems the College Board wants you to know this AP Chemistry study guide gives you winning test taking tips multiple choice strategies and topic guidelines as well as great advice on optimizing your study time and hitting the top of your game on test day This user friendly guide helps you prepare without perspiration by developing a pre test plan organizing your study time and getting the most out of your AP course You ll get help understanding atomic structure and bonding grasping atomic geometry understanding how colliding particles produce states and so much more To provide students with hands on experience AP chemistry courses include extensive labwork as part of the standard curriculum This is why the book dedicates a chapter to providing a brief review of common laboratory equipment and techniques and another to a complete survey of recommended AP chemistry experiments Two full length practice exams help you build your confidence get comfortable with test formats identify your strengths and weaknesses and focus your studies You ll discover how to Create and follow a pretest plan Understand everything you must know about the exam Develop a multiple choice strategy Figure out displacement combustion and acid base reactions Get familiar with stoichiometry Describe patterns and predict properties Get a handle on organic chemistry nomenclature Know your way around laboratory concepts tasks equipment and safety Analyze laboratory data Use practice exams to maximize your score Additionally you ll have a chance to brush up on the math skills that will help you on the exam learn the critical types of chemistry problems and become familiar with the annoying exceptions to chemistry rules Get your own copy of AP Chemistry For Dummies to build your confidence and test taking know how so you can ace that exam

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[Organic Chemistry Study Guide with Solutions Manual](#) Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer

[Elementary Organic Chemistry](#) Homer Adkins, Samuel Marion McElvain, 1928

Descriptive Inorganic Chemistry Geoff Rayner-Canham, Geoffrey Rayner-Canham, Tina Overton, 2009-12-28 This bestselling text introduces descriptive inorganic chemistry in a less rigorous less mathematical way The book uses the periodic table as basis for understanding chemical properties and uncovering relationships between elements in different groups Rayner Canham and Overton s text also familiarizes students with the historical background of inorganic chemistry as well as with its crucial applications especially in regard to industrial processes and environmental issues resulting in a comprehensive appreciation and understanding of the field and the role it will play in their fields of further study

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