



## Teacher Guide: Titration

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### Learning Objectives

- understand the chemical theory underlying titration and how it
- applies the appropriate pH of a substance in a test
- interpret and explain the shape of titration curves
- use an indicator substance that reacts preferentially to react with an analyte
- use titration to determine the concentration of an analyte
- illustrate an appropriate indicator through titrating a weak acid or weak base

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### Materials

acid, alkali, bases, indicators, pipettes, burets, indicator, phenolphthalein, sodium hydroxide, pH, strong acid, strong base, weak acid, weak base, weak acid, weak base

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### Lesson Overview

Students are introduced to the concept of titration, which is a chemical reaction between an acid and a base. The reaction is used to determine the concentration of an acid or base. The reaction is used to determine the concentration of an acid or base. The reaction is used to determine the concentration of an acid or base.

At the end of the lesson, students are asked to write a paragraph explaining the concept of titration. The paragraph should explain the concept of titration and how it is used to determine the concentration of an acid or base.

The following questions are designed to assess student understanding.

- 1. What is the purpose of titration? (To determine the concentration of an acid or base.)
- 2. What is the difference between a strong acid and a weak acid? (A strong acid is completely ionized in water, while a weak acid is only partially ionized.)
- 3. What is the difference between a strong base and a weak base? (A strong base is completely ionized in water, while a weak base is only partially ionized.)



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### Suggested Student Responses

1. The student activity: **Indicator-based titration** (100 minutes)  
Students will work in pairs to perform a titration. They will use a strong acid (hydrochloric acid) and a weak base (sodium hydroxide) to determine the concentration of the weak base. The student will use a pH indicator (phenolphthalein) to determine the endpoint of the titration. The student will record the volume of acid added and the volume of base added. The student will calculate the concentration of the weak base.

# Teacher Guide Titration

**Katrin Zwirgmaier**



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