

# Atomic Structure & Isotope

## Part A: Atomic Structure

1. Complete the following table.

Element name	Atomic number	Mass number	Number of protons	Number of electrons	Number of neutrons	Element symbol
		70				
	48					
				80		
			35			

2. Identify the neutral atom described by name and mass number (example, oxygen-16).

i) The atom with 2 neutrons and 1 proton is \_\_\_\_\_.

ii) The atom with 17 electrons and 18 neutrons is \_\_\_\_\_.

iii) The atom with 6 protons and 8 neutrons is \_\_\_\_\_.

## Part B: Isotope

1. Complete the following table.

Symbol	# of protons	# of neutrons	# of electrons	Symbol	# of protons	# of neutrons	# of electrons
$^1\text{H}_1$				$^{17}\text{O}_8$			
$^2\text{H}_1$				$^{39}\text{K}_{19}$			
$^4\text{He}_2$				$^{40}\text{Ca}_{20}$			
$^6\text{He}_2$				$^{238}\text{U}_{92}$			
$^6\text{Li}_3$				$^{244}\text{Pu}_{94}$			
$^7\text{Li}_3$				$^{56}\text{Fe}_{26}^{3+}$			

# Teaching Transparency Worksheet Isotopes

**McGraw-Hill Staff**



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