

## MOLECULAR GEOMETRY AND POLARITY

### Act. # 1 : Complete Me

Compound	Lewis Structure	Electron Groups	Electron Group Geometry	Bonded Atoms	Lone Pairs	Bond Angle	Molecular Geometry
1. HCN	$\text{H}-\text{C}\equiv\text{N:}$	2	Linear	2	0	$180^\circ$	Linear
2. $\text{NCl}_3$	<pre>       :Cl:         :Cl:-N-Cl:               :Cl:           </pre>	4	Tetrahedral	3	1	$109.5^\circ$	Trigonal bipyramidal
3. $\text{CH}_4$	<pre>       H         H - C - H               H           </pre>	4	Tetrahedral	4	0	$109.5^\circ$	Tetrahedral
4. $\text{PCl}_5$	<pre>       :Cl:         :Cl:-P-Cl:               :Cl:           </pre>	5	Trigonal bipyramidal	5	0	$120^\circ$ $90^\circ$	Trigonal bipyramidal

### Act. # 2 : Identify our bond

	Electronegativity Difference	Type of Bond
1. H—H bond in $\text{H}_2$	$2.1-2.1 = 0.0$	Nonpolar covalent bond
2. Br—Cl bond in $\text{BrCl}_3$	$3.0-2.8 = 0.2$	Nonpolar covalent bond
3. H—Br bond in $\text{HBr}$	$2.8-2.1 = 0.7$	Polar covalent bond
4. As—Cl bond in $\text{AsCl}_3$	$3.0-2.0 = 1.0$	Polar covalent bond
5. P—Br bond in $\text{PBr}_3$	$2.8-2.1 = 0.7$	Polar covalent bond

# Test Molecular Shape Answer Key

**David Baud**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name David Baud.

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