

Chemistry diagnostic test (Topic 12)

1. B – the  $p^4$  config. of S is less stable, electron- electron repulsion.

~~electron-electron repulsion~~  $\uparrow\downarrow \uparrow \uparrow$

2. A – B is lower than Be because two s electrons shield the 2p electrons
3. E – Au and Ag both have full d orbitals(10electrons) ( $s^1$ )
4. E – electronegativity increases from lower left to upper right (F-most electronegative)
5. B – sulfur's oxidation numbers: 2-, 2+, 4+, 6+
6. A – atomic radii increase right to left and top to bottom
7. C – cations are smaller, the higher the charge the smaller the cation
8. A – Group 1 has low 1<sup>st</sup> ionization energy, but 2<sup>nd</sup> is pulling an inner core electron(big jump)
9. E – equal moles - 0.25
10. D – greatest molar mass
11. D – greatest mass/volume
12. E – all same temp
13. E – Phosphorus – high reactivity with oxygen not found free in nature
14. B – not all have a 2+ oxidation
15. A – all substances contain London dispersion forces
16. D – hydrogen must be bonded to nitrogen, oxygen or fluorine
17. A – propane (molar mass – 44) dispersion forces only – lowest temp  
Ethanol (molar mass – 46) also has hydrogen bonding  
Water (molar mass – 18) very strong hydrogen bonds
18. C
19. A – balanced equation is  
 $\text{Na}_2\text{CO}_3 + 2\text{HCl} \rightarrow 2\text{NaCl} + \text{CO}_2 + \text{H}_2\text{O}$   
Limiting is sodium carbonate
20. C –  $(0.300 \text{ moles/L})(.300\text{L}) = 0.0900 \text{ mol Cl}^- \text{ ion} - \text{requires } 1/3 \text{ as much aluminum chloride or } 0.0300 \text{ mol}$   
 $0.0300 \text{ mol}/0.500\text{M or mol/L} = 0.0600\text{L} = 60.0 \text{ ml}$
21. B -  $0.500 \text{ M} \times 0.250\text{L} \times 200\text{g/mol} = 25.0 \text{ g}$
22. D-  $2 \text{ C}_3\text{H}_8\text{O} + 9\text{O}_2 \rightarrow 6\text{CO}_2 + 8\text{H}_2\text{O}$
23. C – 3.00 carbon =  $\frac{1}{2}$  mole of C therefore  $\frac{1}{2}$  mol of Z combines  
 $0.5 \text{ mol Z} \times 36.0 \text{ g/mol}(3\text{timesC}) = 18.0 \text{ g Z}$
24. E – 4 moles of O
25. C – an alkali metal reacts with water to produce hydrogen gas and an aqueous alkali metal hydroxide
26. D –  $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{NaHSO}_4 + \text{H}_2\text{O}$  (diprotic acid and monobasic base)
27. E – X has a 3+ charge
28. E
29. B – 1 mole of magnesium requires 2 moles of HCl therefore HCl is limiting
30. D – a formation reaction produces 1 mole of a product

# Texas Eoc Chemistry Diagnostic Test Answers

**JS Bruner**



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