

Section 2.9

binary compounds
binary ionic compounds
oxanions
binary covalent compounds

Naming of compounds

- Binary compounds
 - Type I—contain a metal that always forms the same cation
 - Type II—contain a metal that can form more than one cation
 - Type III—contain two nonmetals
- Compounds containing a polyatomic ion
- Acids

Discussion Questions

These questions are designed to be considered by groups of students in class. Often these questions work well for introducing a particular topic in class.

- You may have noticed that when water boils, you can see bubbles that rise to the surface of the water. What is inside these bubbles? Explain.
 - air
 - hydrogen and oxygen gas
 - oxygen gas
 - water vapor
 - carbon dioxide gas
- Which of the following is true about an individual atom?
 - An individual atom should be considered a solid.
 - An individual atom should be considered a liquid.
 - An individual atom should be considered a gas.
 - The state of the atom depends on which element it is.
 - An individual atom cannot be considered a solid, liquid, or gas.

Justify your choice, and for those you did not choose, explain why they are incorrect.
- How would you go about finding the number of “chalk molecules” it takes to write your name on the board? Explain what you would need to do and provide a sample calculation.
- These questions concern the work of J. J. Thomson:
 - From what you know of Thomson’s work, which particles do you think he would believe are most important in the formation of compounds (chemical changes) and why?
 - Of the remaining two subatomic particles, which do you place second in importance for forming compounds and why?
 - Propose three models that explain Thomson’s findings and evaluate them. Include Thomson’s findings.
- Heat is applied to an ice cube in a closed container until only steam is present. Draw a representation of this process, assuming you can see it at an extremely high level of magnification. What happens to the size of the molecules? What happens to the total mass of the sample?
- You have a chemical in a sealed glass container filled with air. The setup is sitting on a balance, as shown. The chemical is ignited by means of a magnifying glass focusing sunlight on the reactant. After the chemical has completely burned, which of the following is true? Explain your answer.
 - The balance will read less than 250.0 g.
 - The balance will read 250.0 g.
 - The balance will read greater than 250.0 g.

- Cannot be determined without knowing the identity of the chemical.



- The vitamin niacin (nicotinic acid, $C_6H_5NO_2$) can be isolated from a variety of natural sources such as liver, yeast, milk, and whole grain. It also can be synthesized from commercially available materials. From a nutritional point of view, which source of nicotinic acid is best for use in a multivitamin tablet? Why?
- One of the best indications of a useful theory is that it raises more questions for further experimentation than it originally answered. Is this true of Dalton’s atomic theory? Give examples.
- Dalton assumed that all atoms of the same element are identical in all their properties. Explain why this assumption is not valid.
- How does Dalton’s atomic theory account for each of the following?
 - the law of conservation of mass
 - the law of definite proportion
 - the law of multiple proportions
- What refinements had to be made in Dalton’s atomic theory to account for Gay-Lussac’s results on the combining volumes of gases?
- Which (if any) of the following can be determined by knowing the number of protons in a neutral element? Explain your answer.
 - the number of neutrons in the neutral element
 - the number of electrons in the neutral element
 - the name of the element
- The average mass of a carbon atom is 12.011. Assuming you were able to pick up only one carbon atom, the chance that you would randomly get one with a mass of 12.011 is
 - 0%.
 - 0.011%.
 - about 12%.
 - 12.011%.
 - greater than 50%.
 - None of these is true.

Explain.

Test 46 Chapter 1test Answers

Richard Bailey

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