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 kg the molality equation is a mathematical formula that describes how much a substance weighs in this equation m moles
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 with 1200 ml of 2m hno3 if the final density of solution is same as pure water then molarity of the iron in resulting solution
 which decides the nature of the above solution is 1 0 25m 2 0 5m 3 0 888m 4 none of

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 pressure and temperature changes concentrations expressed in molality are used

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 the molar concentration of the solute is sometimes abbreviated by putting square brackets around the chemical formula of

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signifies solutions labelled with molar concentration a 1.0 M solution contains 1 mole of solute per litre of solution molality is the number of solute moles per kilogram

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web 1 molarity mv mass molar mass $\times 0.4500 \text{ l} = 0.825 \text{ g}$ 141 9579 g/mol $\times 0.0129 \text{ m} = 2$ molality $0.825 \text{ g} / 141.9579 \text{ g/mol} = 0.00581158 \text{ mol/kg}$ 0.00581158 mol/kg $0.4500 \text{ kg} \times 0.0129 \text{ m} = 3$ mole fraction $n_2 / (n_2 + n_1) = 0.825 \text{ g} / 141.9579 \text{ g/mol} = 0.00581158 \text{ mol}$ 2 o 450.0 g 18.015 g/mol 24.97918401 mol

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